

Stoichiometry

Quiz 11*9 Section 9

Section 9: Introduction to Stoichiometry

In the space provided, write the letter of the best response that best completes each problem or best answers each question.

1. Stoichiometry is the study of chemistry that deals with, depending on context, one or more chemical relationships.
 - a. balancing
 - b. identifying
 - c. energy conversion
 - d. mass relationships
 - e. physical measurements
2. Some basic data necessary to develop an atomic hypothesis are:
 - a. the atomic mass, molar mass, and mass
 - b. the atomic gas law
 - c. the atomic mass, molar mass, and measurements
 - d. measurements alone
3. The number of significant figures in a product or quotient is usually the same as the number of significant figures in:
 - a. the number of significant figures in the factor with the least number of significant figures
 - b. the number of significant figures in the factor with the most significant figures
 - c. the number of significant figures of any component given in the problem
 - d. the number of significant figures in the number given in the problem
4. The chemical equation $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$ can be interpreted correctly as:
 - a. 1 molecule of P_4O_{10} + 6 molecules of H_2O form 4 molecules of H_3PO_4
 - b. 1 mole of P_4O_{10} + 6 moles of H_2O form 4 moles of H_3PO_4
 - c. 1 g of P_4O_{10} + 6 g of H_2O form 4 g of H_3PO_4
 - d. both (a) and (b)
5. How many moles of H_2SO_4 are produced from the following chemical equation?
 - a. 1
 - b. 2
 - c. 3
 - d. 4



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Russo, Mike Silver**



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