THE MOLE AND AVOGADRO'S NUMBER

Name _____

One male of a substance contains Avogadro's Number (6.02 x 1025) of molecules

How many molecules are in the quantities below?

- 1. 2.0 moles x = 20 mg (x (6.02 x 13 3 molecules) = 1.2 x 10 4 molecules
- 2. 1.5 moles x = 1.5 mole x (6.02 x 13 molecules) = 9.0 x 10 molecules
- 3. 0.75 mole x=0.75 m/ol x (6.02 x 13 molecular)=4.5 x 10 molecular
- 4. 15 moles (6.02×1023 molecula) = 9.0×104 moleculas
- 5. 0.35 mole $\times (602 \times 10^{23} \text{ molecular}) = 2.1 \times 10^{23} \text{ molecular}$

How many moles are in the number of molecules below?

- 1. 6.02 × 1003 × moles = 6.02 × 133 moleulas (1mol 32) = 1.0 mol.
- 2. 1.204 x 1034 X moles : 1.204 x 134 molecula (1 mol 6.02 x 133 molecula) 2.0 molecular
- 3. 1.5 × 10²⁰

 × moler = 1.5 × 10 molegala (1 mol (0.00024 mole)

 × moler = 1.5 × 10 molegala (1 mol (0.00024 mole)

 = 2.5 × 10⁴ molegala
- 4. 3.4 × 10²⁴

 × moles: 8.4 × 18 moley les (1 mol 602× 10° molecules) = 565 moles
- 5. 7.5 × 1010

 × mole = 7.5 × 15 molearles (1mol (6.02 × 15 molearles) = 1.2 × 10 mole

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Charles Alfred Coulson

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