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# SEXIEST MAN ALIVE

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# Sexiest Man Alive English Edition

**John Mercer, Clarissa Smith**



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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}}$

mole  $\text{H}_2\text{S}$  x 18g  $\text{H}_2\text{O}$  i mole  $\text{H}_2\text{O}$  149 4g  $\text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137}{149} \times 100 = 91.77\%$   
 91 77 problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al x  $\frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times 101.964 \text{ g Al}_2\text{O}_3 = 81.2 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{66.0}{81.2} \times 100 = 81.3\%$

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