

Gas Stoichiometry Worksheet

Name:

Solve all the following gas law problems. Show all work, answers are given at the end of the problem.

Molar Volume

1. Calculate the number of moles contained in 550.mL of carbon dioxide at STP. (0.0246mol)

2. Calculate the mass of 1.50 L of CH_4 at STP. (1.07g)

3. Calculate the volume in liters of 50.0 grams of nitrogen dioxide at STP. (24.3L)

Gas Stoichiometry

4. Given the following reaction: $\text{NH}_4\text{NO}_2 (\text{s}) \rightarrow \text{N}_2 (\text{g}) + 2\text{H}_2\text{O} (\text{g})$ How many liters of nitrogen gas is produced if 50.0L of water is produced at STP? (25.0L)

5. Given the following reaction: $\text{CuO} (\text{s}) + \text{H}_2 (\text{g}) \rightarrow \text{Cu} (\text{s}) + \text{H}_2\text{O} (\text{g})$ If 250. L of hydrogen gas are used to reduce copper (II) oxide at STP, what mass of copper is to be expected? (709g)

6. How many grams of solid lithium must be added to liquid water in order to obtain 15.0 L of hydrogen gas at STP? (9.29g)



Self Check Answers Gas Stoichiometry

Michal Rosen-Zvi

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Self Check Answers Gas Stoichiometry:

Self Check Answers Gas Stoichiometry Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Self Check Answers Gas Stoichiometry**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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