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Section 4.2 The Structure of an Atom

This section compares the properties of these subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Reading Strategy page 1000

Monitoring Your Understanding Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Resources Handbook** at the end of your textbook.

What I Know About Me	What I Would Like To Learn	What I Have Learned

Properties of Subatomic Particles pages 188-197

- What are three subatomic particles?
a. _____ b. _____ c. _____
- Circle the letter that identifies a subatomic particle with a positive charge.
a. nucleus b. proton
c. neutron d. electron
- Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? _____

Comparing Subatomic Particles pages 135-140

- Circle the letters of properties that vary among subatomic particles.
 - color
 - mass
 - charge
 - location in the atom
- Circle the letter of the expression that accurately compares the masses of neutrons and protons.
 - mass of 1 neutron = mass of 1 proton
 - mass of 2000 neutrons = mass of 1 proton
 - mass of 1 electron = mass of 1 proton
 - mass of 1 neutron = mass of 1 electron

Section Structure Of The Atom Answer

KJ Lindholm-Leary



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