



12. Number of atoms in 560 g of Fe (atomic mass 56 g mol^{-1}) is :
 (1) Twice that in 70 g N (2) Half that in 20 g H
 (3) Both (1) and (2) (4) None of these
13. Which has maximum number of atoms :
 (1) 24 g of C (12) (2) 56 g of Fe (56) (3) 27 g of Al (27) (4) 108 g Ag (108)
14. The total number of protons, electrons and neutrons in 12 g of $^{12}_6\text{C}$ is :
 (1) 1.084×10^{25} (2) 6.022×10^{23} (3) 6.022×10^{22} (4) 18
15. 1.0 g of hydrogen contains 6×10^{23} atoms. The atomic weight of helium is 4. It follows that the number of atoms in 1 g of He is :
 (1) $\frac{1}{4} \times 6 \times 10^{23}$ (2) $4 \times 6 \times 10^{23}$ (3) 6×10^{23} (4) 12×10^{23}
16. A sample of aluminium has a mass of 54.0 g. What is the mass of the same number of magnesium atoms? (At. wt. Al = 27, Mg = 24)
 (1) 12 g (2) 24 g (3) 48 g (4) 96 g.
17. The number of atoms in 558.5 g of Fe (at wt. = 55.85) is :
 (1) Twice that in 60 g carbon (2) 6.022×10^{23}
 (3) Half in 8 g He (4) $558.5 \times 6.023 \times 10^{23}$
18. The total number of electrons present in 8.0 g of methane is.
 (1) 4.8×10^{24} (2) 3.01×10^{24} (3) 4.8×10^{23} (4) 3.01×10^{23} .
19. How many moles of electron weigh one kilogram :
 (1) 6.023×10^{23} (2) $\frac{1}{9.108} \times 10^{31}$ (3) $\frac{6.023}{9.108} \times 10^{24}$ (4) $\frac{1}{9.108 \times 6.023} \times 10^6$
20. Under the same conditions, two gases have the same number of molecules. They must
 (1) be noble gases (2) have equal volumes
 (3) have a volume of 22.4 dm^3 each (4) have an equal number of atoms
21. 5.6 litre of oxygen at STP contains :
 (1) 6.02×10^{23} atoms (2) 3.01×10^{23} atoms
 (3) 1.505×10^{23} atoms (4) 0.7525×10^{23} atoms
22. If V ml of the vapours of substance at NTP weight W g. Then molecular weight of substance is :
 (1) $(W/V) \times 22400$ (2) $\frac{V}{W} \times 22.4$ (3) $(W-V) \times 22400$ (4) $\frac{W \times 1}{V \times 22400}$
23. If 1 dm^3 of a gas weights 2.5 g at STP, its gram-molecular weight is :
 (1) 56 g (2) 11.2 g (3) 22.4 g (4) 224 g
24. 16 g of an ideal gas SO_x occupies 5.6 L. at STP. The value of x is
 (1) $x = 3$ (2) $x = 2$ (3) $x = 4$ (4) none of these

Section (C) : Density and Vapour density

1. Vapour density of a gas if its density is 0.178 g/L at NTP is :
 (1) 0.178 (2) 2 (3) 4 (4) 0.089
2. A gas is found to have the formula $(\text{CO})_x$. It's VD is 70 the value of x must be :
 (1) 7 (2) 4 (3) 5 (4) 6
3. Number of electrons in 1.8 mL of $\text{H}_2\text{O}(l)$ is about :
 (1) 6.02×10^{23} (2) 3.011×10^{23} (3) 0.6022×10^{21} (4) 60.22×10^{20}

Section (D) : Percentage composition and Molecular formula

1. The percentage value of nitrogen in urea is about
 (1) 46 (2) 85 (3) 18 (4) 28

Section Chemistry Answers Molecular Composition

Eugene Gustave Eberle



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