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Acids and Bases Reinforcement Activity

Write "A" on the line if the following substance is an acid. Write "B" on the line if it is a base. If the substance is generally not an acid or a base, write "N" for neither. Most bases are metal hydroxides (OFI), metal oxides (OF), or armonium (NH₄") hydroxides/oxides.

L H ₂ O	15. Ca(OH) _k	29. HCi
z. H ₂ SO ₄		30. Li ₂ O
s. NivOH	17. HNOs	31. HBr
4. MgO	18. CeHoOs	32. NapO
s. HCN	19. HNOs	33. HI
S. HCIO4	20. HCx0x0x	34. HF
7. HCKO _{3.}	21. PO ₄ *	35. K ₂ O
k HCIO ₁	22 Mg ^{3*}	36. H ₃ S
9. SOV2	23. H ₂ C ₂ O ₄	37. C ₂ H ₄
IO. HCIO	24 H ₂ BO ₃	38. C ₁ H ₄
II. NieOH	25. HiSOs	39. KMeOv
12. LiOH	26. SrO	40. HxO0v
13. Bu(OH)s		41. HsPOs
14. NHAOH	28 Rb ₁ O	42. SerOH)s

Complete and balance the acid-base neutralization reactions. Write "s" under the acid (reactant), and write "b" under the base (reactant). Write "s" under the salt, which is the new ionic compound produced. Hint: carbonate ions (CO₃³) will split spart to produce CO₂ gas and to help from water.

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