



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a. Na_2O (ionic) b. CH_4 (covalent) c. HCl (covalent)
 d. KNO_3 (ionic) e. H_2O (covalent)

2. Write the formula for the following molecular compounds.

- a. carbon dioxide b. sulfur dioxide c. carbon monoxide
 d. N_2O e. SO_2 f. CO

3. Write the Lewis structure for each of the following molecular compounds.

a. CH_4 (carbon has 4 valence electrons and hydrogen has 1; there are 10 valence electrons in total and 4 bonds must be formed in the molecule)

b. H_2O (oxygen has 6 valence electrons and hydrogen has 1; there are 8 valence electrons in total and 2 bonds must be formed in the molecule)

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .

4. Draw the Lewis structure for acetylene, C_2H_2 .



5. Draw the Lewis structure for ammonia, NH_3 . Be sure to show all lone pairs of electrons.



(Lone pairs are shown on the nitrogen atom.)

SECTION 8.3 BONDING: THE ORBITAL

1. Write the hybridization for each carbon in the following molecules: CH_4 , C_2H_2 . The degree of hybridization is the number of hybrid orbitals.

2. Write the hybridization for each carbon in the following: C_2H_4 . (Hybridization is the number of hybrid orbitals.)

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