



COVALENT BONDING

Practice Problems

Use your knowledge to answer the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a. Na_2O (ionic) b. CH_4 (covalent) c. HCl (covalent)
 d. KNO_3 (ionic) e. H_2O (covalent)

2. Write the formula for the following molecular compounds.

- a. nitrogen triiodide b. phosphorus pentachloride c. carbon
 d. sulfur e. sulfur

3. Write the formula for the following molecular compounds: nitrogen triiodide.

4. Write the formula for the following molecular compounds: phosphorus pentachloride.

5. Write the formula for the following molecular compounds: carbon tetrachloride.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .



4. Draw the Lewis structure for acetylene, C_2H_2 .

5. Draw the Lewis structure for the following: $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$.



6. Draw the Lewis structure for the following: $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$.

SECTION 8.3 BONDING: THE ORBITAL

1. Write the Lewis structure for the following: $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$.

2. Write the Lewis structure for the following: $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$.

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J Elliott



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