

Forming Chemical Bonds

Section 8.1

- I. Chemical Bonds
 - a. The force that holds two atoms together is called a chemical bond.
 - b. Chemical bonds may form by the attraction between a positive nucleus and negative electrons or the attraction between a positive ion and a negative ion.
 - c. Valence electrons are involved in the formation of chemical bonds between two atoms.
 - d. How do thousands of compounds form from the relatively few elements known to exist?
 - i. The answer lies in the electron structure of the atoms of the elements involved and the nature of the attractive forces between these atoms.
 - ii. Recall that electron-dot structure is a type of diagram used to keep track of valence electrons and is especially useful when illustrating the formation of chemical bonds.
 - iii. Recall that ionization energy refers to how easily an atom loses an electron. The term electron affinity indicates how much attraction an atom has for electrons.
 - iv. The difference in reactivity is directly related to the valence electrons.
- II. Formation of Positive Ions
 - a. Compare sodium and neon electron configuration.
 - b. Compare sodium ion and neutral sodium.
 - c. It is possible for atoms to achieve a *pseudo-noble gas configuration*.
 - i. Zinc electron configuration?
 1. What electrons would it lose?
 - d. A positively charged ion is called a **cation**.
- III. Formation of Negative Ions
 - a. Recall nonmetals have a great attraction for electrons and form a stable outer configuration by gaining electrons.
 - b. Chlorine electron configuration?
 - i. What happens?
 - ii. What noble gas does it become like?
 - c. A negatively charged ion is called an **anion**.

The Formation and Nature of Ionic Bonds

Section 8.2

- I. Formation of an Ionic Bond
 - a. The electrostatic force that holds oppositely charged particles together in an ionic compound is referred to as an **ionic bond**.
- II. Compounds that contain ionic bonds are ionic compounds.
 - a. If ionic bonds occur between metals and nonmetal oxygen, oxides form.
 - b. Most other ionic compounds are called salts.
 - c. Binary ionic compounds contain a metallic cation and nonmetallic anion.
- III. What happens in the formation of a compound?
 - a. Explain the formation of the ionic compound composed of each pair of elements.
 - i. sodium and nitrogen

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