



COVALENT BONDING

Practice Problems

Do your best to do better than your previous problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Draw a molecular formula for each of the following compounds.

- a. Silicon tetrachloride b. Nitrogen dioxide c. Hydrogen peroxide
d. Sulfur dioxide e. Dinitrogen pentoxide

2. Write the formula for the following molecular compounds.

- a. Carbon monoxide b. Diphosphorus pentoxide c. Nitrogen
d. Sulfur e. Hydrogen f. Oxygen

3. Draw a Lewis structure for each of the following molecular compounds.

a. Diphosphorus pentoxide b. Carbon dioxide c. Nitrogen dioxide d. Hydrogen peroxide e. Hydrogen sulfide f. Hydrogen cyanide g. Hydrogen chloride h. Hydrogen bromide i. Hydrogen iodide j. Hydrogen fluoride

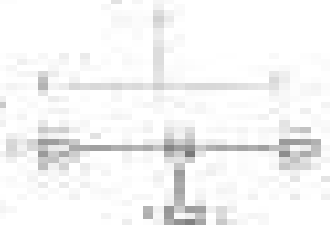
4. Draw a Lewis structure for each of the following molecular compounds. Label each atom as being an electrophile, nucleophile, or both.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw a Lewis structure for each of the following compounds.



2. Draw a Lewis structure for each of the following compounds.



3. Draw a Lewis structure for each of the following compounds.



4. Draw a Lewis structure for each of the following compounds.

5. Draw a Lewis structure for each of the following compounds. Label each atom as being an electrophile, nucleophile, or both.



6. Draw a Lewis structure for each of the following compounds.

SECTION 8.3 BONDING: THE ORBITAL

1. Draw a Lewis structure for each of the following compounds. Label each atom as being an electrophile, nucleophile, or both.

2. Draw a Lewis structure for each of the following compounds. Label each atom as being an electrophile, nucleophile, or both.

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