

Chapter 6 Chemical Bonds

Section 6.2 Covalent Bonding**(pages 166–169)**

This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules.

Reading Strategy **page 166**

Relating Text and Visuals As you read the section, look closely at Figure 5. Complete the table for describing each type of model shown. For more information on this Reading Strategy, see the **Reading and Study Skills** in the Skills and Reference Handbook at the end of your textbook.

Molecular Models	
Model	Description
Electron dot	
Structural formula	
Space-filling	
Electron cloud	

Covalent Bonds **(pages 166–169)**

- Describe a covalent bond. _____
- Circle the letters of molecular models that show orbitals of atoms overlapping when a covalent bond forms.
 - electron dot
 - structural formula
 - space-filling
 - electron cloud
- Describe a molecule. _____
- Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. _____
- Circle the correct answer. Nitrogen has five valence electrons. How many pairs of electrons must two nitrogen atoms share in order for each atom to have eight valence electrons?
 - zero
 - one
 - two
 - three

Section 6 2 Covalent Bonding Answer Key

M Tight



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