

Chapter 6 Chemical Bonds

Section 6.1 Ionic Bonding

(pages 155–164)

This section describes the formation of ionic bonds and the properties of ionic compounds.

Reading Strategy (page 155)

Sequencing As you read, complete the concept map to show what happens to atoms during ionic bonding. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



Stable Electron Configurations (page 155)

- Describe the type of electron configuration that makes an atom stable and not likely to react. Atoms with a full outer shell of electrons are stable and not likely to react.
- Describe an electron dot diagram. An electron dot diagram is a model in which each atom is represented by a symbol, and the valence electrons are represented by dots around the symbol.

Ionic Bonds (pages 155–161)

- Some elements achieve stable electron configurations through the transfer of electrons between atoms.
- By losing one valence electron, a sodium atom achieves the same electron arrangement as an atom of neon.
- Circle the letter that states the result of a sodium atom transferring an electron to a chlorine atom.
 - Each atom ends up with a more stable electron arrangement.
 - The sodium atom becomes more stable, but the chlorine atom becomes less stable.
 - The chlorine atom becomes more stable, but the sodium atom becomes less stable.
 - Each atom ends up with a less stable electron arrangement.
- Is the following sentence true or false? An ion is an atom that has a net positive or negative electric charge. True
- An ion with a negative charge is called an anion.

Section 6 1 Ionic Bonding Worksheet

Karin Nielsen-Saines

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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Table of Contents Section 6 1 Ionic Bonding Worksheet

1. Understanding the eBook Section 6 1 Ionic Bonding Worksheet
 - The Rise of Digital Reading Section 6 1 Ionic Bonding Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 6 1 Ionic Bonding Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 6 1 Ionic Bonding Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 6 1 Ionic Bonding Worksheet
 - Personalized Recommendations
 - Section 6 1 Ionic Bonding Worksheet User Reviews and Ratings
 - Section 6 1 Ionic Bonding Worksheet and Bestseller Lists
5. Accessing Section 6 1 Ionic Bonding Worksheet Free and Paid eBooks
 - Section 6 1 Ionic Bonding Worksheet Public Domain eBooks
 - Section 6 1 Ionic Bonding Worksheet eBook Subscription Services
 - Section 6 1 Ionic Bonding Worksheet Budget-Friendly Options

6. Navigating Section 6 1 Ionic Bonding Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 6 1 Ionic Bonding Worksheet Compatibility with Devices
 - Section 6 1 Ionic Bonding Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 6 1 Ionic Bonding Worksheet
 - Highlighting and Note-Taking Section 6 1 Ionic Bonding Worksheet
 - Interactive Elements Section 6 1 Ionic Bonding Worksheet
8. Staying Engaged with Section 6 1 Ionic Bonding Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 6 1 Ionic Bonding Worksheet
9. Balancing eBooks and Physical Books Section 6 1 Ionic Bonding Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 6 1 Ionic Bonding Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section 6 1 Ionic Bonding Worksheet
 - Setting Reading Goals Section 6 1 Ionic Bonding Worksheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 6 1 Ionic Bonding Worksheet
 - Fact-Checking eBook Content of Section 6 1 Ionic Bonding Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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