

CHAPTER 8 REVIEW

Chemical Equations and Reactions**SECTION 1: Describing Chemical reactions**

1. Match the symbol on the left with its appropriate description on the right.

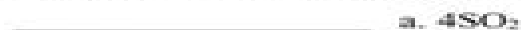


- (a) A precipitate forms.
(b) A gas forms.
(c) A reversible reaction occurs.
(d) Heat is applied to the reactants.
(e) A chemical is dissolved in water.
(f) A chemical is in the liquid state.

2. Finish balancing the following equation:



3. In each of the following formulas, write the total number of atoms present.



4. Convert the following word equation into a balanced chemical equation: aluminum metal + copper(II) fluoride
- $\rightarrow$
- aluminum fluoride + copper metal

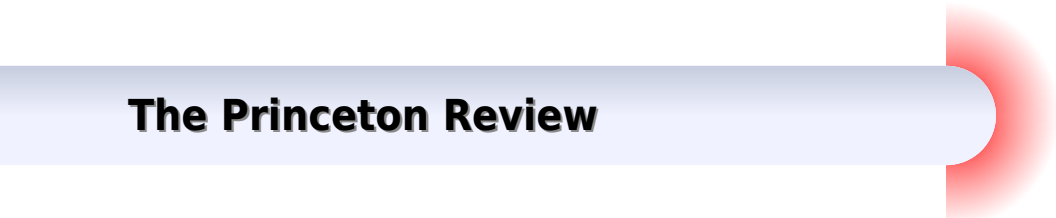


5. One way to test the salinity of a water sample is to add a few drops of silver nitrate solution with a known concentration. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, and sodium nitrate is left in solution. Translate these sentences into a balanced chemical equation.



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