

Name _____ Period _____

Acids and Bases Reinforcement Activity

Write "A" on the line if the following substance is an acid. Write "B" on the line if it is a base. If the substance is generally not an acid or a base, write "N" for neither. Most bases are metal hydroxides (OH), metal oxides (O²⁻), or ammonium (NH₄⁺) hydroxides/oxides.

1. H₂O _____

2. H₂SO₄ _____

3. NaOH _____

4. MgO _____

5. HCN _____

6. HClO₄ _____

7. HClO₃ _____

8. HClO₂ _____

9. SO₃²⁻ _____

10. HClO _____

11. NaOH _____

12. LiOH _____

13. Ba(OH)₂ _____

14. NH₄OH _____

15. Ca(OH)₂ _____

16. CaO _____

17. HNO₃ _____

18. C₆H₁₂O₆ _____

19. HNO₂ _____

20. HC₉O₈O₅ _____

21. PO₄³⁻ _____

22. Mg²⁺ _____

23. H₃C₇O₈ _____

24. H₃BO₃ _____

25. H₂SO₃ _____

26. SrO _____

27. RbOH _____

28. Rb₂O _____

29. HCl _____

30. Li₂O _____

31. HBr _____

32. NaO _____

33. HI _____

34. HF _____

35. K₂O _____

36. H₂S _____

37. C₃H₈ _____

38. C₃H₆ _____

39. KMnO₄ _____

40. H₂CO₃ _____

41. H₃PO₄ _____

42. Sr(OH)₂ _____

Complete and balance the acid-base neutralization reactions. Write "a" under the acid (reactant), and write "b" under the base (reactant). Write "c" under the salt, which is the new ionic compound produced. Hint: carbonate ions (CO₃²⁻) will split apart to produce CO₂ gas and to help form water.



Section 3 Reinforcement Acids Bases And Salts

Jianjun Gao



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