

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

Na, K, S, O, Cl

3. Which atom has a higher electronegativity: fluorine or chlorine?

Fluorine is more electronegative than chlorine because it is smaller.

4. Which two elements in the same family generally have similar properties?

There have been many cases where atoms in the same family have similar properties.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (decreasing to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest atomic radius? Which element has the smallest atomic radius? Which element has the greatest electronegativity?

Period 7, Francium, Helium, Fluorine, Astatine

7. Which trend in ionization energy values across a period is expected and the periodic table? Which element has the highest ionization energy? Which element has the lowest ionization energy?

Ionization energy increases from left to right across a period.

8. Write the atoms in each group that have the largest atomic radius.

a. **Li**
 b. **Na**
 c. **K**
 d. **Rb**
 e. **Cs**
 f. **Fr**

a. **Li**
 b. **Na**
 c. **K**
 d. **Rb**
 e. **Cs**
 f. **Fr**

9. Write the atoms in each group that have the smallest atomic radius.

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Clifford Lane



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