13 Al -> "e + 26 Ma 48K -> = e + 40 G balance nuclear reaction equations aluminum-26 undergoes positron emission potassium-40 undergoes beta decay (electron) 238 () -> 4 He + 234 Th uranium-238 undergoes alpha decay electron configurations 152 25226 35236 451345 Write the electron configuration for C. Write the electron configuration for Cr. Write the shorthand notation for Bi. De] 65244145 110603 NOMENCLATURE [naming compounds, writing formulas] Write the formula for each of the following compounds: a. Sulfur hexafluoride SFS b. Lithium nitride Link c. Chromium (III) carbonate Cr2 ((0))3 d. Tin (II) chloride Sn Ce2 e. Ammonium acetate NHUCZHOR (OF CHOCOONHY) f. Mercury (I) chloride Hazelz g. Potassium bromate KBrog h. Hydrosulfuric acid H2 Some i. Chloric acid Sulfurous acid 42300 Name each of the following compounds. a Cuso, copper (II) sulfate aluminum flooride b. AIF. c. HI d. NO narosen monovale c. H₂Sc (40) hydroselen ic acid f. HNO, notric acid narous acrd g. HNO: NaHSO_a Soulden

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g? 40 -> 20 -> 10 -> 5 -> 2.5 -> 1.25 ->

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

3 half lives 100-750-25-3 [2.5 mg]

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65,

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

63
$$\times .692 = 43.596 \rightarrow 49$$

65 $\times .309 = 20.02 \rightarrow 20.$
 $\frac{+}{63.62} \rightarrow 64$

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Mary Louise Turgeon

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Table of Contents Section 7 1 Review Answers Chemistry

- 1. Understanding the eBook Section 7 1 Review Answers Chemistry
 - The Rise of Digital Reading Section 7 1 Review Answers Chemistry
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Section 7 1 Review Answers Chemistry
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 7 1 Review Answers Chemistry
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Section 7 1 Review Answers Chemistry
 - Personalized Recommendations
 - Section 7 1 Review Answers Chemistry User Reviews and Ratings
 - Section 7 1 Review Answers Chemistry and Bestseller Lists

- 5. Accessing Section 7 1 Review Answers Chemistry Free and Paid eBooks
 - Section 7 1 Review Answers Chemistry Public Domain eBooks
 - Section 7 1 Review Answers Chemistry eBook Subscription Services
 - Section 7 1 Review Answers Chemistry Budget-Friendly Options
- 6. Navigating Section 7 1 Review Answers Chemistry eBook Formats
 - o ePub, PDF, MOBI, and More
 - Section 7 1 Review Answers Chemistry Compatibility with Devices
 - Section 7 1 Review Answers Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 7 1 Review Answers Chemistry
 - Highlighting and Note-Taking Section 7 1 Review Answers Chemistry
 - Interactive Elements Section 7 1 Review Answers Chemistry
- 8. Staying Engaged with Section 7 1 Review Answers Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 7 1 Review Answers Chemistry
- 9. Balancing eBooks and Physical Books Section 7 1 Review Answers Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 7 1 Review Answers Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 7 1 Review Answers Chemistry
 - Setting Reading Goals Section 7 1 Review Answers Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 7 1 Review Answers Chemistry
 - Fact-Checking eBook Content of Section 7 1 Review Answers Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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