

Acids and Bases

Reinforcement Activity

Write "A" on the line if the following substance is an acid. Write "B" on the line if it is a base. If the substance is generally not an acid or a base, write "N" for neither. Most bases are metal hydroxides (OH), metal oxides (O²⁻), or ammonium (NH₄⁺) hydroxides/oxides.

- | | | |
|---|---|--|
| 1. H ₂ O _____ | 15. Ca(OH) ₂ _____ | 29. HCl _____ |
| 2. H ₂ SO ₄ _____ | 16. CaO _____ | 30. Li ₂ O _____ |
| 3. NaOH _____ | 17. HNO ₃ _____ | 31. HBr _____ |
| 4. MgO _____ | 18. C ₆ H ₁₂ O ₆ _____ | 32. NaO _____ |
| 5. HCN _____ | 19. HNO ₂ _____ | 33. HI _____ |
| 6. HClO ₄ _____ | 20. HC ₉ O ₈ O ₅ _____ | 34. HF _____ |
| 7. HClO ₃ _____ | 21. PO ₄ ³⁻ _____ | 35. K ₂ O _____ |
| 8. HClO ₂ _____ | 22. Mg ²⁺ _____ | 36. H ₂ S _____ |
| 9. SO ₄ ²⁻ _____ | 23. H ₃ C ₇ O ₈ _____ | 37. C ₃ H ₈ _____ |
| 10. HClO _____ | 24. H ₃ BO ₃ _____ | 38. C ₃ H ₈ _____ |
| 11. NaOH _____ | 25. H ₂ SO ₃ _____ | 39. KMnO ₄ _____ |
| 12. LiOH _____ | 26. SrO _____ | 40. H ₂ CO ₃ _____ |
| 13. Ba(OH) ₂ _____ | 27. KOH _____ | 41. H ₃ PO ₄ _____ |
| 14. NH ₄ OH _____ | 28. Rb ₂ O _____ | 42. Sr(OH) ₂ _____ |

Complete and balance the acid-base neutralization reactions. Write "a" under the acid (reactant), and write "b" under the base (reactant). Write "c" under the salt, which is the new ionic compound produced. Hint: carbonate ions (CO₃²⁻) will split apart to produce CO₂ gas and to help form water.

1. _____ NaOH + _____ HCl →

2. _____ HC₃H₇O₃ + _____ Li₂CO₃ →

3. _____ H₂SO₄ + _____ Al(OH)₃ →

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S Nieto



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Table of Contents Section Acids And Bases Reinforcement Answer Key

1. Understanding the eBook Section Acids And Bases Reinforcement Answer Key
 - The Rise of Digital Reading Section Acids And Bases Reinforcement Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Section Acids And Bases Reinforcement Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Acids And Bases Reinforcement Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section Acids And Bases Reinforcement Answer Key
 - Personalized Recommendations
 - Section Acids And Bases Reinforcement Answer Key User Reviews and Ratings
 - Section Acids And Bases Reinforcement Answer Key and Bestseller Lists
5. Accessing Section Acids And Bases Reinforcement Answer Key Free and Paid eBooks
 - Section Acids And Bases Reinforcement Answer Key Public Domain eBooks
 - Section Acids And Bases Reinforcement Answer Key eBook Subscription Services
 - Section Acids And Bases Reinforcement Answer Key Budget-Friendly Options
6. Navigating Section Acids And Bases Reinforcement Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Section Acids And Bases Reinforcement Answer Key Compatibility with Devices
 - Section Acids And Bases Reinforcement Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section Acids And Bases Reinforcement Answer Key
 - Highlighting and Note-Taking Section Acids And Bases Reinforcement Answer Key
 - Interactive Elements Section Acids And Bases Reinforcement Answer Key
 8. Staying Engaged with Section Acids And Bases Reinforcement Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Acids And Bases Reinforcement Answer Key
 9. Balancing eBooks and Physical Books Section Acids And Bases Reinforcement Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section Acids And Bases Reinforcement Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Section Acids And Bases Reinforcement Answer Key
 - Setting Reading Goals Section Acids And Bases Reinforcement Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Section Acids And Bases Reinforcement Answer Key
 - Fact-Checking eBook Content of Section Acids And Bases Reinforcement Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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