



# Polk Audio Db65manual

**WJ Hussar**



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mol ko 2 and 0 mol h<sub>2</sub> o what is the limiting reactant

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu}_2\text{AgNO}_3 + 2\text{AgNO}_3$$

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web a write a balanced equation for the reaction  $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g  $\text{Al}$   $\times \frac{1 \text{ mol Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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 80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  x 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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

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