

## Chapter 5 The Periodic Table

## Section 5.2 The Modern Periodic Table

(pages 130–138)

This section explains the organization of the modern periodic table and discusses the general properties of metals, nonmetals, and metalloids.

## Reading Strategy (page 130)

**Previewing** Before you read, complete the table by writing two questions about the periodic table on pages 132–133. As you read, write answers to your questions. For more information on this reading strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

| Questions About the Periodic Table                                        |                                                                                                                              |
|---------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------|
| Question                                                                  | Answer                                                                                                                       |
| Students might ask:<br>What does atomic mass mean?                        | Atomic mass is a value that depends on the distribution of an element's isotopes in nature and the masses of those isotopes. |
| Why are two columns of elements placed before the main body of the table? | Two kinds of elements are placed before the table to make the table more compact.                                            |

Other possible questions: Why are there two numbering systems for the columns? Why is period 7 incomplete?

## The Periodic Law (pages 131–133)

1. Is the following sentence true or false? In the modern periodic table, elements are arranged by increasing number of protons.  
\_\_\_\_\_ **true** \_\_\_\_\_
2. Explain why the number of elements per period varies. \_\_\_\_\_  
\_\_\_\_\_ **The number of elements per period varies because the number of available orbitals increases from energy level to energy level.** \_\_\_\_\_
3. Properties of elements repeat in a predictable way when atomic numbers are used to arrange elements into groups. This pattern of repeating properties is called the \_\_\_\_\_  
\_\_\_\_\_ **periodic law** \_\_\_\_\_

## Atomic Mass (page 134)

4. Label the four types of information supplied for chlorine in the diagram.

|    |          |
|----|----------|
| a. | 17       |
| b. | Cl       |
| c. | Chlorine |
| d. | 35.453   |

- Atomic number**
- Element symbol**
- Element name**
- Atomic mass**

# Periodic Table Section 5 2 Answer

**Atef Korchef**



## Periodic Table Section 5 2 Answer:

Student Solutions Manual for Physical Chemistry C. A. Trapp, Peter Atkins, Julio de Paula, 2009-12-18 With its modern emphasis on the molecular view of physical chemistry its wealth of contemporary applications vivid full color presentation and dynamic new media tools the thoroughly revised new edition is again the most modern most effective full length textbook available for the physical chemistry classroom Available in Split Volumes For maximum flexibility in your physical chemistry course this text is now offered as a traditional text or in two volumes Volume 1 Thermodynamics and Kinetics ISBN 1 4292 3127 0 Volume 2 Quantum Chemistry Spectroscopy and Statistical Thermodynamics ISBN 1 4292 3126 2 **Chemistry All-in-One For Dummies (+ Chapter Quizzes Online)** Christopher R. Hren, John T. Moore, Peter J. Mikulecky, 2022-11-23 Everything you need to crush chemistry with confidence Chemistry All in One For Dummies arms you with all the no nonsense how to content you ll need to pass your chemistry class with flying colors You ll find tons of practical examples and practice problems and you ll get access to an online quiz for every chapter Reinforce the concepts you learn in the classroom and beef up your understanding of all the chemistry topics covered in the standard curriculum Prepping for the AP Chemistry exam Dummies has your back with plenty of review before test day With clear definitions concise explanations and plenty of helpful information on everything from matter and molecules to moles and measurements Chemistry All in One For Dummies is a one stop resource for chem students of all valences Review all the topics covered in a full year high school chemistry course or one semester of college chemistry Understand atoms molecules and the periodic table of elements Master chemical equations solutions and states of matter Complete practice problems and end of chapter quizzes online Chemistry All In One For Dummies is perfect for students who need help with coursework or want to cram extra hard to ace that chem test

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