

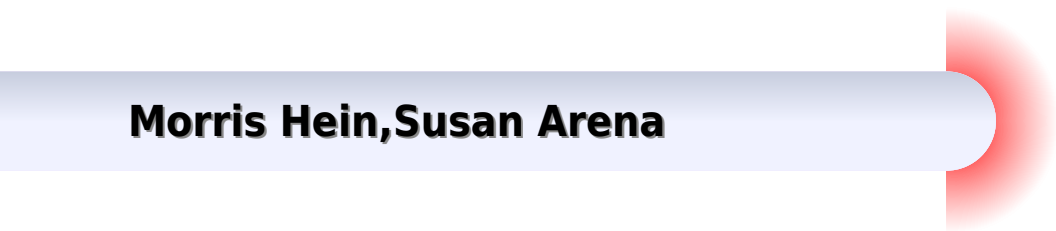
Name _____

Development of Modern Atomic Theory – Video worksheet

1. The development of atomic theory started over 2000 years ago with a man named _____.
2. He believed that all matter is made up of tiny particles called _____.
3. Who was the next person to advance atomic theory in 1808? _____
4. He thought that atoms of a given element are identical in _____, _____, and other properties. Today we know this to NOT be true because of _____.
5. 90 years later, a man named J.J. Thompson discovered the _____ in 1897.
6. His model of the atom is referred to as the _____ model.
7. A Crooke's tube is essentially the same as a _____ tube.
8. Ernest Rutherford won the nobel prize in 1908 for his discovery of the _____.
9. His famous experiment is known as the _____ experiment.
10. Rutherford's model of the atom is known as the _____ model of the atom.
11. What are the tiny particles in the nucleus of the atom? _____
12. Who was the Danish physicist who won the nobel prize in 1922? _____
13. In his model, electrons move around the nucleus of the atom in a _____ motion.
14. In the Bohr model, a _____ is emitted with energy $E = hf$ when an electron jumps down from one orbit to another.
15. Louis de Broglie was a French physicist who made groundbreaking contributions to _____ theory.
16. He postulated that all matter exhibits _____ properties.
17. Our current understanding of the atom is based on the _____ Wave Mechanical Model of the Atom.
18. The Heisenberg Uncertainty Principle states that no experiment can measure both the _____ and _____ of a quantum particle simultaneously.
19. The Schrodinger wave equation shows that atomic particles like _____ exhibit wave-like behaviors.
20. The wave mechanical model of the atom shows the _____ of finding an electron in an atom.

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Morris Hein, Susan Arena



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