

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the chemical equation
- 2. Explain how to interpret the equation
- 3. List the steps for writing a chemical equation

Vocabulary

- 1. chemical equation
- 2. chemical equation
- 3. reactant
- 4. coefficient
- 5. balanced equation

Part A Completion

Write the coefficient for each in the chemical equation, and balance the chemical equation. Write the name of each reactant and product.

1. A chemical equation is always represented by a balanced _____ 1 _____
 The chemical equation is a chemical change called _____ 2 _____ that uses
 chemical substances (reactants) to make _____ 3 _____ (products) and energy.
 The law of conservation of _____ 4 _____ is always followed. Matter cannot be
 created or destroyed. The _____ 5 _____ of the reactants must be
 equal to the _____ 6 _____ of the products. The _____ 7 _____ of the
 atoms must be the same on both sides of the equation.

2. The chemical equation is always balanced by a balanced _____ 8 _____
 coefficient. The coefficient is _____ 9 _____ that a value _____ 10 _____ is put in front
 of _____ 11 _____ substance. The coefficient is always _____ 12 _____ of 1. _____ 13 _____
 must be placed in front of a compound to show a change in the number of molecules of the
 compound.

- 1. _____
- 2. _____
- 3. _____
- 4. _____
- 5. _____
- 6. _____
- 7. _____
- 8. _____
- 9. _____
- 10. _____
- 11. _____
- 12. _____
- 13. _____

Section Review 1describing Chemical Reactions

M Lipman



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