

9

CHEMICAL NAMES AND FORMULAS

SECTION 9.1 NAMING IONS (pages 253–258)

This section explains the use of the periodic table to determine the charge of an ion. It also defines polyatomic ion and gives the names and formulas for the most common polyatomic ions.

► Monatomic Ions (pages 253–256)

1. What are monatomic ions?

Monatomic ions are ions consisting of only one atom.

2. How is the ionic charge of a Group 1A, 2A, or 3A ion determined?

The ionic charge is numerically equal to the group number.

3. How is the ionic charge of a Group 5A, 6A, or 7A ion determined?

The charge of an ion in Groups 5A, 6A, or 7A is determined by subtracting 8 from the group number.

4. Circle the letter of the type of element that often has more than one common ionic charge.

- a. alkali metal
- b. alkaline earth metal
- ☒ c. transition metal
- d. nonmetal

5. The Stock system of naming transition metal cations uses a Roman numeral in parentheses to indicate the numeric value of the ionic charge.

6. An older naming system uses the suffix *-ous* to name the cation with the lesser charge, and the suffix *-ic* to name the cation with the greater charge.

7. What is a major advantage of the Stock system over the old naming system?

The Stock system gives the actual charge of the ion.

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Christian Drosten



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Table of Contents Section Quiz Chemical Names And Formulas Key

1. Understanding the eBook Section Quiz Chemical Names And Formulas Key
 - The Rise of Digital Reading Section Quiz Chemical Names And Formulas Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Section Quiz Chemical Names And Formulas Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Quiz Chemical Names And Formulas Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section Quiz Chemical Names And Formulas Key
 - Personalized Recommendations
 - Section Quiz Chemical Names And Formulas Key User Reviews and Ratings
 - Section Quiz Chemical Names And Formulas Key and Bestseller Lists
5. Accessing Section Quiz Chemical Names And Formulas Key Free and Paid eBooks

- Section Quiz Chemical Names And Formulas Key Public Domain eBooks
 - Section Quiz Chemical Names And Formulas Key eBook Subscription Services
 - Section Quiz Chemical Names And Formulas Key Budget-Friendly Options
6. Navigating Section Quiz Chemical Names And Formulas Key eBook Formats
- ePub, PDF, MOBI, and More
 - Section Quiz Chemical Names And Formulas Key Compatibility with Devices
 - Section Quiz Chemical Names And Formulas Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Section Quiz Chemical Names And Formulas Key
 - Highlighting and Note-Taking Section Quiz Chemical Names And Formulas Key
 - Interactive Elements Section Quiz Chemical Names And Formulas Key
8. Staying Engaged with Section Quiz Chemical Names And Formulas Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Quiz Chemical Names And Formulas Key
9. Balancing eBooks and Physical Books Section Quiz Chemical Names And Formulas Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section Quiz Chemical Names And Formulas Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section Quiz Chemical Names And Formulas Key
- Setting Reading Goals Section Quiz Chemical Names And Formulas Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section Quiz Chemical Names And Formulas Key
- Fact-Checking eBook Content of Section Quiz Chemical Names And Formulas Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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