

Name: Answer Key Date: \_\_\_\_\_ Per.: \_\_\_\_\_

# Covalent Bond Practice

1. Fill in the missing information on the chart.

| Element    | # of Protons | # of Electrons | # of Valence Electrons | # of electrons to fill outer shell. |
|------------|--------------|----------------|------------------------|-------------------------------------|
| Carbon     | 6            | 6              | 4                      | 4                                   |
| Hydrogen   | 1            | 1              | 1                      | 1                                   |
| Chlorine   | 17           | 17             | 7                      | 1                                   |
| Helium     | 2            | 2              | 2                      | 0                                   |
| Phosphorus | 15           | 15             | 5                      | 3                                   |
| Oxygen     | 8            | 8              | 6                      | 2                                   |
| Sulfur     | 16           | 16             | 6                      | 2                                   |
| Nitrogen   | 7            | 7              | 5                      | 3                                   |

2. For each of the following covalent bonds:

- Write the symbols for each element.
- Draw a Lewis Dot structure for the valence shell of each element.
- Rearrange the electrons to pair up electrons from each atom.
- Draw circles to show the sharing of electrons between each pair of atoms.
- Draw the bond structure using chemical symbols and lines. Use one line for each pair of electrons that is shared.
- Write the chemical formula for each molecule.

a) Hydrogen + Hydrogen



b) Chlorine + Chlorine



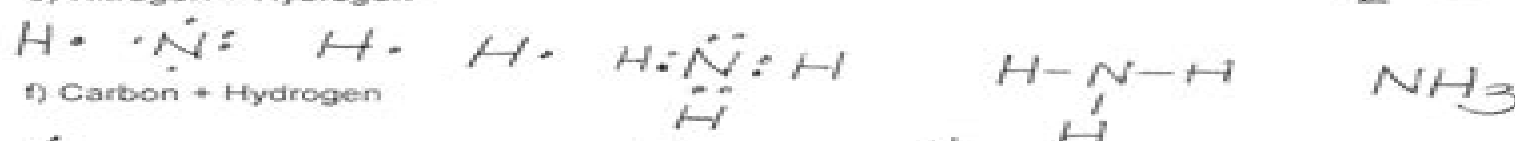
c) Hydrogen + Chlorine



d) Hydrogen + Oxygen



e) Nitrogen + Hydrogen



f) Carbon + Hydrogen



## Section 8 The Covalent Bond Answer Key

**Wenbin Ji**



## Section 8 The Covalent Bond Answer Key:

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