

Concepts in Chemistry

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev's original short periodic table gave useful chemical properties of elements appeared in regular intervals and showed elements were arranged because of not knowing
 - (a) density
 - (b) electronegativity
 - (c) atomic number
 - (d) atomic mass
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a new period
 - (c) adding a new column
 - (d) adding a new group
5. _____ The noble gases that property of the noble gases is that they are
 - (a) unstable
 - (b) inert
 - (c) radioactive
 - (d) highly reactive
6. _____ Likewise, the first column in Group 1, has characteristic number of 1. The second column in the group has characteristic number of
 - (a) 2
 - (b) 10
 - (c) 18
 - (d) 32
7. How groups in the periodic table are numbered? (a) and (b) atomic number of 10
 - (a) 1 to 18
 - (b) 1 to 10
 - (c) 1 to 18
 - (d) 1 to 10
8. How many periods are there in the periodic table?
 - (a) 18
 - (b) 10
 - (c) 18
 - (d) 10
9. How many elements are in the periodic table?
 - (a) 118
 - (b) 118
 - (c) 118
 - (d) 118
10. How many elements are in the periodic table?
 - (a) 118
 - (b) 118
 - (c) 118
 - (d) 118

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Clemens Wendtner



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