

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete chemical equation.

Vocabulary

- 1. chemical equation
- 2. chemical equation
- 3. reactant
- 4. coefficient
- 5. balanced equation

Part A Completion

Write the appropriate word(s) in space or under the blank line, and underline the word(s) in the correct definition. Write complete Latin and Greek names in italics.

A chemical equation is usually represented by a symbol _____. The chemical formula for a chemical compound is _____. The law of conservation of mass states that the _____ of matter remains the same throughout a reaction. A chemical equation is written with _____ on the left and _____ on the right. _____ is the symbol for oxygen. _____ is the symbol for hydrogen. _____ is the symbol for carbon. _____ is the symbol for nitrogen. _____ is the symbol for sulfur. _____ is the symbol for phosphorus. _____ is the symbol for chlorine. _____ is the symbol for bromine. _____ is the symbol for iodine. _____ is the symbol for fluorine. _____ is the symbol for neon. _____ is the symbol for argon. _____ is the symbol for krypton. _____ is the symbol for xenon. _____ is the symbol for radon.

Chemical equations are written with _____ on the left and _____ on the right. The _____ of matter is conserved in a chemical reaction. _____ is the symbol for oxygen. _____ is the symbol for hydrogen. _____ is the symbol for carbon. _____ is the symbol for nitrogen. _____ is the symbol for sulfur. _____ is the symbol for phosphorus. _____ is the symbol for chlorine. _____ is the symbol for bromine. _____ is the symbol for iodine. _____ is the symbol for fluorine. _____ is the symbol for neon. _____ is the symbol for argon. _____ is the symbol for krypton. _____ is the symbol for xenon. _____ is the symbol for radon.

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