

8 • Electron Configurations & Periodicity

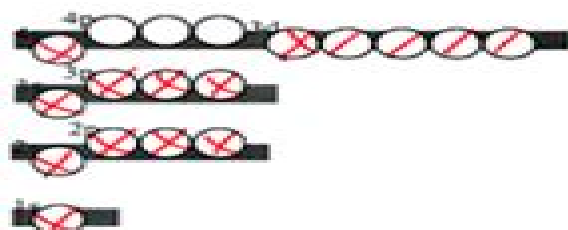
WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ s: <u>18</u>	Element: Mg # of e ⁻ s: <u>12</u>	Element: N # of e ⁻ s: <u>7</u>	Element: Li # of e ⁻ s: <u>3</u>	Element: P # of e ⁻ s: <u>15</u>	Element: Cl # of e ⁻ s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar $1s^2 2s^2 2p^6 3s^2 3p^6$
 Ar $[\text{Ne}] 3s^2 3p^6$
- Mg $1s^2 2s^2 2p^6 3s^2$
 Mg $[\text{Ne}] 3s^2$
- N $1s^2 2s^2 2p^3$
 N $[\text{He}] 2s^2 2p^3$
- Li $1s^2 2s^1$
 Li $[\text{He}] 2s^1$
- P $1s^2 2s^2 2p^6 3s^2 3p^3$
 P $[\text{Ne}] 3s^2 3p^3$
- Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
 Cl $[\text{Ne}] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



Section 5 3 Electron Configurations Answer Key

S Ben Porath



Section 5 3 Electron Configurations Answer Key:

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Table of Contents Section 5 3 Electron Configurations Answer Key

1. Understanding the eBook Section 5 3 Electron Configurations Answer Key
 - The Rise of Digital Reading Section 5 3 Electron Configurations Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 5 3 Electron Configurations Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 5 3 Electron Configurations Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 5 3 Electron Configurations Answer Key
 - Personalized Recommendations
 - Section 5 3 Electron Configurations Answer Key User Reviews and Ratings
 - Section 5 3 Electron Configurations Answer Key and Bestseller Lists
5. Accessing Section 5 3 Electron Configurations Answer Key Free and Paid eBooks
 - Section 5 3 Electron Configurations Answer Key Public Domain eBooks
 - Section 5 3 Electron Configurations Answer Key eBook Subscription Services

- Section 5 3 Electron Configurations Answer Key Budget-Friendly Options
- 6. Navigating Section 5 3 Electron Configurations Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 5 3 Electron Configurations Answer Key Compatibility with Devices
 - Section 5 3 Electron Configurations Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 5 3 Electron Configurations Answer Key
 - Highlighting and Note-Taking Section 5 3 Electron Configurations Answer Key
 - Interactive Elements Section 5 3 Electron Configurations Answer Key
- 8. Staying Engaged with Section 5 3 Electron Configurations Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 5 3 Electron Configurations Answer Key
- 9. Balancing eBooks and Physical Books Section 5 3 Electron Configurations Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 5 3 Electron Configurations Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 5 3 Electron Configurations Answer Key
 - Setting Reading Goals Section 5 3 Electron Configurations Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 5 3 Electron Configurations Answer Key
 - Fact-Checking eBook Content of Section 5 3 Electron Configurations Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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