

Wintersemester 2019/2020 - Klausuraufgaben 8B

Aufgabe 8 (max. 70 Minuten/100 Punkte)

Die nachstehenden Probleme sollen Sie zur Vertiefung des wissenschaftlichen Verständnisses der Stoffeigenschaften (Kapitel 10) befähigen. Sie sind, soweit in der Aufgabenstellung nicht anders angegeben, unter folgenden Bedingungen zu lösen:

$$p = p_0 = 1 \text{ bar} = 10^5 \text{ Pa}$$

1. Chloroform (Molare Masse $M = 119,38 \text{ g/mol}$) wird bei einer Temperatur von 25°C verdampft. Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle?

$$\bar{v} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \cdot 8,314 \text{ J/(mol}\cdot\text{K)} \cdot 298,15 \text{ K}}{0,11938 \text{ kg/mol}}} = 308,4 \text{ m/s}$$

2. Ethylacetat (Molare Masse $M = 88,10 \text{ g/mol}$) wird bei 25°C verdampft. Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle?

$$\bar{v} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \cdot 8,314 \text{ J/(mol}\cdot\text{K)} \cdot 298,15 \text{ K}}{0,08810 \text{ kg/mol}}} = 319,4 \text{ m/s}$$

3. Methanol ist ein verdampfendes flüchtiges Gas. Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle bei einer Temperatur von 25°C ? Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle bei einer Temperatur von 50°C ?

$$\bar{v} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \cdot 8,314 \text{ J/(mol}\cdot\text{K)} \cdot 298,15 \text{ K}}{0,03204 \text{ kg/mol}}} = 447,1 \text{ m/s}$$

4. Die mittlere Geschwindigkeit $\bar{v}$ der Moleküle von H_2 beträgt bei einer bestimmten Temperatur $1,7 \text{ km/s}$. Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle von O_2 bei derselben Temperatur? Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle von H_2 bei einer Temperatur, bei der die mittlere Geschwindigkeit $\bar{v}$ der Moleküle von O_2 $1,7 \text{ km/s}$ beträgt?

$$\bar{v} = \sqrt{\frac{3RT}{M}} \Rightarrow \bar{v} \propto \frac{1}{\sqrt{M}} \Rightarrow \bar{v}_{\text{H}_2} = \bar{v}_{\text{O}_2} \cdot \sqrt{\frac{M_{\text{O}_2}}{M_{\text{H}_2}}} = 1,7 \text{ km/s} \cdot \sqrt{\frac{32}{2}} = 5,26 \text{ km/s}$$

5. Ethylacetat wird bei einer bestimmten Temperatur verdampft. Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle bei einer Temperatur von 25°C ? Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle bei einer Temperatur von 50°C ?

$$\bar{v} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \cdot 8,314 \text{ J/(mol}\cdot\text{K)} \cdot 298,15 \text{ K}}{0,08810 \text{ kg/mol}}} = 319,4 \text{ m/s}$$

6. Ethylacetat wird bei einer bestimmten Temperatur verdampft. Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle bei einer Temperatur von 25°C ? Wie groß ist die mittlere Geschwindigkeit $\bar{v}$ der Moleküle bei einer Temperatur von 50°C ?

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