

## Polyatomic Ions WS

Ammonium ( $\text{NH}_4$ )<sup>1+</sup>  
 Carbonate ( $\text{CO}_3$ )<sup>2-</sup>  
 Chromate ( $\text{CrO}_4$ )<sup>2-</sup>

Dichromate ( $\text{Cr}_2\text{O}_7$ )<sup>2-</sup>  
 Hydroxide ( $\text{OH}$ )<sup>1-</sup>  
 Nitrate ( $\text{NO}_3$ )<sup>1-</sup>

Nitrite ( $\text{NO}_2$ )<sup>1-</sup>  
 Sulfate ( $\text{SO}_4$ )<sup>2-</sup>  
 Sulfite ( $\text{SO}_3$ )<sup>2-</sup>

Phosphate ( $\text{PO}_4$ )<sup>3-</sup>

**Write the NAME of each of the following compounds.**

1.  $(\text{NH}_4)\text{Cl}$  Ammonium Chloride
2.  $\text{Be}(\text{SO}_4)$  Beryllium Sulfate
3.  $(\text{NH}_4)_3\text{N}$  Ammonium Nitride
4.  $\text{MgCl}_2$  Magnesium Chloride
5.  $\text{NH}_4(\text{NO}_3)$  Ammonium Nitrate
6.  $\text{Sr}_3(\text{PO}_4)_2$  Strontium Phosphate
7.  $\text{Zn}(\text{CrO}_4)_2$  Zinc Dichromate
8.  $\text{K}_2(\text{Cr}_2\text{O}_7)$  Potassium Dichromate

9.  $\text{Ga}(\text{ClO}_3)_3$  Gallium Chlorate
10.  $\text{Cu}(\text{OH})$  Copper I Hydroxide
11.  $(\text{NH}_4)_3(\text{PO}_4)$  Ammonium Phosphate
12.  $\text{Fe}(\text{SO}_4)$  Iron II Sulfate
13.  $\text{Mg}(\text{NO}_2)_2$  Magnesium Nitrite
14.  $(\text{NH}_4)\text{NO}_2$  Ammonium Nitrite
15.  $\text{Na}_2(\text{Cr}_2\text{O}_7)$  Sodium Dichromate
16.  $\text{Na}(\text{OH})$  Sodium Hydroxide

**Write the CHEMICAL FORMULA for each of the given NAMES.....cross charges**

- |                         |                                     |   |                               |                   |
|-------------------------|-------------------------------------|---|-------------------------------|-------------------|
| 17. calcium carbonate   | $\text{Ca}^{2+} \text{CO}_3^{2-}$   | = | $\text{Ca}_2(\text{CO}_3)_2$  | = $\text{CaCO}_3$ |
| 18. barium nitrate      | $\text{Ba}^{+2} \text{NO}_3^{-1}$   | = | $\text{Ba}(\text{NO}_3)$      |                   |
| 19. ammonium sulfate    | $\text{NH}_4^{+1} \text{SO}_4^{-2}$ | = | $(\text{NH}_4)_2 \text{SO}_4$ |                   |
| 20. aluminum hydroxide  | $\text{Al}^{+3} \text{OH}^{-1}$     | = | $\text{Al}(\text{OH})_3$      |                   |
| 21. calcium phosphate   | $\text{Ca}^{+2} \text{PO}_4^{-3}$   | = | $\text{Ca}_3(\text{PO}_4)_2$  |                   |
| 22. cesium nitrate      | $\text{Cs} \text{NO}_2^{-1}$        | = | $\text{CsNO}_3$               |                   |
| 23. sodium nitrite      | $\text{Na}^{+1} \text{NO}_3^{-1}$   | = | $\text{NaNO}_3$               |                   |
| 24. calcium sulfate     | $\text{Ca}^{+2} \text{SO}_4^{-2}$   | = | $\text{CaSO}_4$               |                   |
| 25. beryllium sulfate   | $\text{Be} \text{SO}_4^{-2}$        | = | $\text{BeSO}_4$               |                   |
| 26. sodium carbonate    | $\text{Na}^{+1} \text{CO}_3^{-2}$   | = | $\text{Na}_2\text{CO}_3$      |                   |
| 27. magnesium phosphate | $\text{Mg}^{+2} \text{PO}_4^{-3}$   | = | $\text{Mg}_3(\text{PO}_4)_2$  |                   |
| 28. calcium phosphate   | $\text{Ca}^{+2} \text{PO}_4^{-3}$   | = | $\text{Ca}_3(\text{PO}_4)_2$  |                   |

# Polyatomic Ions Ws

**E Durkheim**



## **Polyatomic Ions Ws:**

## **Polyatomic Ions Ws** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has be more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Polyatomic Ions Ws**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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### **Polyatomic Ions Ws Introduction**

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