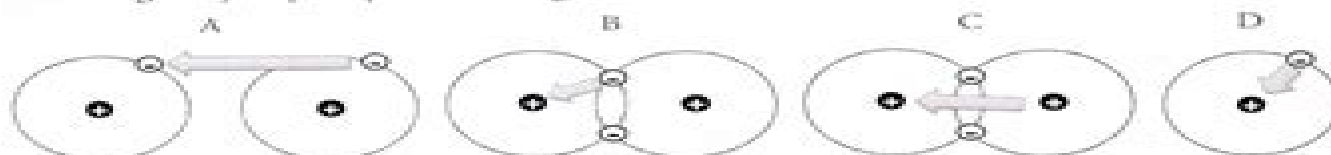


Read This!

Electronegativity is a measure of the ability of an atom's nucleus to attract electrons from a different atom within a covalent bond. A higher electronegativity value correlates to a stronger pull on the electrons in a bond. This value is only theoretical. It cannot be directly measured in the lab.

12. Using the definition stated in the *Read This!* box above, select the best visual representation for electronegativity. Explain your reasoning.



13. Locate the electronegativity values in Model 1.

- a. What is the trend in electronegativity going down a group in Model 1?

Electronegativity decreases as you go down a group (vertical column).

- b. Explain the existence of the trend described in part a in terms of atomic structure and Coulombic attraction.

The ability to attract electrons is based upon attractive force. The greater the distance between the nucleus and valence electrons, the less the attractive force. Less force means a lower electronegativity.

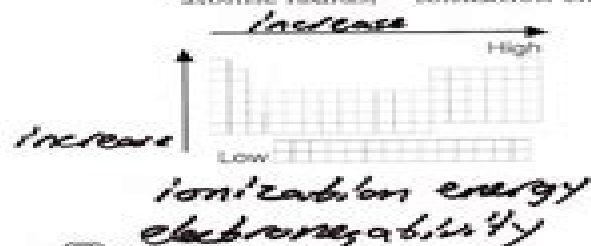
- c. What is the trend in electronegativity going across a period in Model 1?

Electronegativity increases from left to right across a period (horizontal row)

- d. Explain the existence of the trend described in part c in terms of atomic structure and Coulombic attraction.

The greater the number protons in an atom, in a particular period, the greater the attractive force between the nucleus and valence electrons. The greater attractive force means a higher electronegativity.

14. The two diagrams below can summarize each of the three trends discussed in this activity. Write "atomic radius," "ionization energy," and "electronegativity" under the appropriate diagram.



Periodic Trends Pogil Answer Key

Ying Liu



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Table of Contents Periodic Trends Pogil Answer Key

1. Understanding the eBook Periodic Trends Pogil Answer Key
 - The Rise of Digital Reading Periodic Trends Pogil Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Pogil Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Pogil Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Pogil Answer Key
 - Personalized Recommendations
 - Periodic Trends Pogil Answer Key User Reviews and Ratings

- Periodic Trends Pogil Answer Key and Bestseller Lists
- 5. Accessing Periodic Trends Pogil Answer Key Free and Paid eBooks
 - Periodic Trends Pogil Answer Key Public Domain eBooks
 - Periodic Trends Pogil Answer Key eBook Subscription Services
 - Periodic Trends Pogil Answer Key Budget-Friendly Options
- 6. Navigating Periodic Trends Pogil Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Pogil Answer Key Compatibility with Devices
 - Periodic Trends Pogil Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Pogil Answer Key
 - Highlighting and Note-Taking Periodic Trends Pogil Answer Key
 - Interactive Elements Periodic Trends Pogil Answer Key
- 8. Staying Engaged with Periodic Trends Pogil Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Pogil Answer Key
- 9. Balancing eBooks and Physical Books Periodic Trends Pogil Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Pogil Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Pogil Answer Key
 - Setting Reading Goals Periodic Trends Pogil Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Pogil Answer Key
 - Fact-Checking eBook Content of Periodic Trends Pogil Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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