

The change in the number of molecules of one substance is a multiple of the change in the number of molecules of another.

$$\text{rate} = \frac{\Delta[\text{H}_2]}{\Delta t} = -\frac{\Delta[\text{Cl}_2]}{\Delta t} = +\frac{\Delta[\text{HCl}]}{2\Delta t} \quad \text{coefficient}$$

- For the above reaction, for every 1 mol of H_2 used, 1 mol of I_2 will also be used and 2 mol of HI made.
- Therefore, the change in the concentration of each substance is multiplied by 1/coefficient.

For coefficient
1 over coefficient



$$\text{rate} = -\frac{1}{a} \frac{\Delta[A]}{\Delta t} = -\frac{1}{b} \frac{\Delta[B]}{\Delta t} = +\frac{1}{c} \frac{\Delta[C]}{\Delta t} = +\frac{1}{d} \frac{\Delta[D]}{\Delta t}$$

$\nearrow a$
coefficient

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Practice Problem: Reaction Rates

Consider this balanced chemical equation: $\text{rate} = -\frac{1}{3} \frac{\Delta[\text{I}^-]}{\Delta t}$

Quizlet Chemistry Chapter 14

IM Harris



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