

POUMSUNG SCORING SYSTEM (C:\PMS\TKD\POUMSUNG\POUMSUNG.MDI)

BITMATT, PAUL      Koryu Taekwondo Centre   Male Block 1st Dan 31-50      Geumgang

Event Name: City West 2009 Pooms - R      Match Number: #10001 - H      Pattern No: 1 2      Judges: 1 2 3

Category	Sub-Category	Score Allocation									
Movement Accuracy	Accuracy in Basic Movement	0.1 Deductions					0.5 Deductions				
	Accuracy in Individual movement of the Poomsae	-	2	+						-	0
Mastery	Extent of Movement	1	.9	.8	.7	.6	.5	0.9			
	Balance	1	.9	.8	.7	.6	.5	0.9			
	Power and Speed	1	.9	.8	.7	.6	.5	0.9			
Presentation	Co-ordination of Rhythm Tempo / Tone and Power	1	.9	.8	.7	.6	.5	0.9			
	Expression of the 'KI'	1	.9	.8	.7	.6	.5	0			

Match = #10001    Pattern = 2    Judges = 3    < Prev    Next >

# Poomsae Competition Scoring Guidelines

**Hongru Du**

A red circular graphic with a gradient, appearing as a stylized arrow or a decorative element, positioned to the right of the name Hongru Du.

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## **Table of Contents Poomsae Competition Scoring Guidelines**

1. Understanding the eBook Poomsae Competition Scoring Guidelines
  - The Rise of Digital Reading Poomsae Competition Scoring Guidelines
  - Advantages of eBooks Over Traditional Books
2. Identifying Poomsae Competition Scoring Guidelines
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Poomsae Competition Scoring Guidelines
  - User-Friendly Interface
4. Exploring eBook Recommendations from Poomsae Competition Scoring Guidelines
  - Personalized Recommendations
  - Poomsae Competition Scoring Guidelines User Reviews and Ratings
  - Poomsae Competition Scoring Guidelines and Bestseller Lists
5. Accessing Poomsae Competition Scoring Guidelines Free and Paid eBooks
  - Poomsae Competition Scoring Guidelines Public Domain eBooks
  - Poomsae Competition Scoring Guidelines eBook Subscription Services
  - Poomsae Competition Scoring Guidelines Budget-Friendly Options
6. Navigating Poomsae Competition Scoring Guidelines eBook Formats

- ePub, PDF, MOBI, and More
- Poomsae Competition Scoring Guidelines Compatibility with Devices
- Poomsae Competition Scoring Guidelines Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Poomsae Competition Scoring Guidelines
  - Highlighting and Note-Taking Poomsae Competition Scoring Guidelines
  - Interactive Elements Poomsae Competition Scoring Guidelines
- 8. Staying Engaged with Poomsae Competition Scoring Guidelines
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Poomsae Competition Scoring Guidelines
- 9. Balancing eBooks and Physical Books Poomsae Competition Scoring Guidelines
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Poomsae Competition Scoring Guidelines
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Poomsae Competition Scoring Guidelines
  - Setting Reading Goals Poomsae Competition Scoring Guidelines
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Poomsae Competition Scoring Guidelines
  - Fact-Checking eBook Content of Poomsae Competition Scoring Guidelines
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web donor whereas a base generically rep resented by b is a proton acceptor when acids and bases combine they swap a proton ha b a hb strong acids e g sulfuric acid have a very pronounced ten dency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from ph data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral ph at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web this problem has been solved you ll get a detailed solution from a subject matter expert that helps you learn core concepts see answer question part b acid base titration lab report i determining the concentration of sulfuric acid solution  $\text{H}_2\text{SO}_4$  by titrating it with  $\text{NaOH}$  solution

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of  $\text{HCl}$  because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then  $\text{ph} = -\log 0.100 = 1.000$  ph log 0.100 = 1.000 b titrant volume 12.50 ml

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