

Periodic Trends Worksheet

Directions: Use the table to determine the following properties:

1. Write the following elements by increasing atomic radius: sodium, sulfur, aluminum, oxygen, potassium.

K, Na, Al, S, O

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, carbon, aluminum.

Na, Al, S, C, O

3. Which atom has the higher first ionization energy: Na or K?

Na has a higher first ionization energy than K because it is smaller.

4. Which two elements in the same family generally have similar properties?

There have been two atoms that share a same valence shell, potassium and sodium.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Period 7, francium, cesium, barium, potassium, rubidium, and sodium.

7. Which trend in ionization energy values across periods and the periodic table? Which element has the highest?

Increases across the periodic table from sodium, francium, and potassium.

8. Write the atoms in each group that have the highest atomic radius.

Li	Be
Na	Mg
K	Ca
Rb	Sr
Cs	Ba
Fr	Ra

Al	Si	P
Ga	Ge	As
In	Sn	Sb
Tl	Pb	Bi

9. Circle the atom in each group that has the greatest electronegativity.

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