

Bonding and Reactions

Study Guide

10 pages including KEY

Chemical Bonding and Reactions Study Guide

This study packet will help you study for the Chemical Bonding and Reactions exam by highlighting the most important concepts on the exam. This test will evaluate your understanding of the material you have learned in this course and will be used for the purpose of determining your placement in future courses.

B. VOCABULARY

Atom	Electron	Electron transfer	Electron transfer
Atomic mass	Electron	Electron transfer	Electron transfer
Atomic number	Electron	Electron transfer	Electron transfer
Chemical bond	Electron	Electron transfer	Electron transfer
Chemical reaction	Electron	Electron transfer	Electron transfer

C. READING

1. Why do atoms bond together?
2. What is a covalent bond?
3. What kind of electrons form covalent bonds?
4. What is an ionic bond?
5. What kind of electrons form ionic bonds?
6. Why don't noble gases bond together with other elements?

Write the formula for the following compounds with multiple valence numbers.

1. Copper (II) chloride
2. Copper (II) chloride

Write the following compounds with multiple valence numbers.

1. Iron
2. Iron

REACTIONS

What happens in an oxidation-reduction reaction?

What does an oxidation-reduction reaction involve?

What is an oxidation-reduction reaction?

What is an oxidation-reduction reaction?

What is the law of conservation of mass?

What is the law of conservation of mass?

A coefficient is a number placed in front of a chemical formula to indicate the number of molecules of that substance that are involved in the reaction. It is used to balance the equation.

What is a chemical reaction? A chemical reaction is a process in which one or more substances are converted into one or more different substances.

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B. VOCABULARY

1. Atom
2. Atomic mass
3. Atomic number
4. Chemical bond
5. Chemical reaction
6. Coefficient
7. Electron
8. Electron transfer
9. Electron transfer
10. Electron transfer
11. Electron transfer
12. Electron transfer
13. Electron transfer
14. Electron transfer
15. Electron transfer
16. Electron transfer
17. Electron transfer
18. Electron transfer
19. Electron transfer
20. Electron transfer

Bonding and Chemical Reactions - KEY

C. READING

1. Why do atoms bond together?
 - Atoms bond together to achieve a more stable electronic configuration, typically by filling or emptying their valence electron shells to satisfy the octet rule.
2. What is a covalent bond?
 - A covalent bond is a type of chemical bond formed between two non-metal atoms by sharing their valence electrons. This sharing results in a stable electronic configuration for both atoms.
3. What kind of electrons form covalent bonds?
 - Valence electrons, which are the outermost electrons of an atom, form covalent bonds by being shared between the two bonding atoms.
4. What is an ionic bond?
 - An ionic bond is a type of chemical bond formed between a metal atom and a non-metal atom. It is formed by the transfer of one or more valence electrons from the metal to the non-metal, resulting in the formation of oppositely charged ions that are held together by electrostatic attraction.
5. Why don't noble gases bond together with other elements?
 - Noble gases have a full valence shell, which makes them chemically stable and unreactive. They do not have a tendency to gain or lose electrons to achieve a more stable configuration.
6. What is a chemical reaction?
 - A chemical reaction is a process in which one or more substances are converted into one or more different substances. This process involves the breaking and forming of chemical bonds.
7. What is a chemical reaction?
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