

Blackline Master 5.8

Ionic Compounds: Names and Formulas Worksheet

1. Write the formulas for the following compounds.

(a) magnesium oxide	<u>MgO</u>
(b) sodium fluoride	<u>NaF</u>
(c) aluminum nitride	<u>AlN</u>
(d) potassium sulfide	<u>K₂S</u>
(e) lithium iodide	<u>LiI</u>
(f) calcium bromide	<u>CaBr₂</u>
(g) beryllium oxide	<u>BeO</u>
(h) nickel ^(II) chloride	<u>NiCl₂</u>
(i) magnesium nitride	<u>Mg₃N₂</u>
(j) aluminum sulfide	<u>Al₂S₃</u>

(k) copper(I) bromide	<u>CuBr</u>
(l) tin(II) iodide	<u>SnI₂</u>
(m) iron(III) chloride	<u>FeCl₃</u>
(n) calcium phosphide	<u>Ca₃P₂</u>
(o) lead(II) oxide	<u>PbO</u>
(p) lead(IV) fluoride	<u>PbF₄</u>
(q) tin(IV) bromide	<u>SnBr₄</u>
(r) copper(II) sulfide	<u>CuS</u>
(s) iron(II) oxide	<u>FeO</u>
(t) calcium nitride	<u>Ca₃N₂</u>

2. Write the names for the following compounds.

(a) Li ₂ O	<u>lithium oxide</u>
(b) AlCl ₃	<u>aluminum chloride</u>
(c) MgS	<u>magnesium sulfide</u>
(d) CaO	<u>calcium oxide</u>
(e) KBr	<u>potassium bromide</u>
(f) BeF	<u>beryllium fluoride</u>
(g) Na ₃ N	<u>sodium nitride</u>
(h) Al ₂ O ₃	<u>aluminum oxide</u>
(i) CuCl ₂	<u>copper(II) chloride</u>
(j) FeBr ₃	<u>iron(III) bromide</u>

(k) PbS	<u>lead (II) sulfide</u>
(l) SnO ₂	<u>tin (IV) oxide</u>
(m) Na ₂ S	<u>sodium sulfide</u>
(n) Mg ₃ P ₂	<u>magnesium phosphide</u>
(o) NiO	<u>nickel (II) oxide</u>
(p) CuI	<u>copper (I) iodide</u>
(q) PbCl ₂	<u>lead (II) chloride</u>
(r) FeP	<u>iron (III) phosphide</u>
(s) CaF ₂	<u>calcium fluoride</u>
(t) K ₃ P	<u>potassium phosphide</u>

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Table of Contents Physical Science Names And Formulas Worksheet 1

1. Understanding the eBook Physical Science Names And Formulas Worksheet 1
 - The Rise of Digital Reading Physical Science Names And Formulas Worksheet 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Names And Formulas Worksheet 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Names And Formulas Worksheet 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Names And Formulas Worksheet 1

- Personalized Recommendations
- Physical Science Names And Formulas Worksheet 1 User Reviews and Ratings
- Physical Science Names And Formulas Worksheet 1 and Bestseller Lists
- 5. Accessing Physical Science Names And Formulas Worksheet 1 Free and Paid eBooks
 - Physical Science Names And Formulas Worksheet 1 Public Domain eBooks
 - Physical Science Names And Formulas Worksheet 1 eBook Subscription Services
 - Physical Science Names And Formulas Worksheet 1 Budget-Friendly Options
- 6. Navigating Physical Science Names And Formulas Worksheet 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Names And Formulas Worksheet 1 Compatibility with Devices
 - Physical Science Names And Formulas Worksheet 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Names And Formulas Worksheet 1
 - Highlighting and Note-Taking Physical Science Names And Formulas Worksheet 1
 - Interactive Elements Physical Science Names And Formulas Worksheet 1
- 8. Staying Engaged with Physical Science Names And Formulas Worksheet 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Names And Formulas Worksheet 1
- 9. Balancing eBooks and Physical Books Physical Science Names And Formulas Worksheet 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Names And Formulas Worksheet 1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Names And Formulas Worksheet 1
 - Setting Reading Goals Physical Science Names And Formulas Worksheet 1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Names And Formulas Worksheet 1

- Fact-Checking eBook Content of Physical Science Names And Formulas Worksheet 1
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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