

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? (anion/cation) _____
which cation it gains and electron it loses from it? _____ (anion/cation)
2. Are placed in the halogen group has _____ valence electrons, and needs to _____ 1 more to have it. This is because it has _____ valence.
3. Are _____ bond to the adjacent elements oppositely charged ions. (PCl₃ is covalent bond is _____)
4. Is a covalent bond, shared _____ electrons, achieved by the bond bond, all over _____ electrons.
5. (Electron group) Which side from left shared completely by it completely bond is to _____, and which group is shared equally to the _____ bond group it is covalent. In ionic, gain electrons, or negative charged by, both ending with _____
6. Are ionic ions have up to _____ valence electrons. This easily combined with the particular bond that all have 8 valence electrons is called _____
7. Shared with each other, but not shared _____ by another group of atoms.

B. Formulas for Ionic compounds - Put in the right order

Write correct chemical formula	Aluminum	Fluorine	Phosphorus	Chlorine
COVALENT BOND				
IONIC BOND				

3. Write in the chemical formulas for each of the following ions: _____

Write your own ionic bond structures, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

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