

Chapter 3: Periodic Trends (NB p.29-30)

Vocabulary: Period, group, atomic radius, ionization energy.
Equations & Constants: N/A

Practice Problems:

- D 1. The energy required to remove an electron from an atom is the atom's
a. electron affinity.
b. electron energy.
c. electronegativity.
d. ionization energy.

- D 2. The element that has the greatest electronegativity is
a. oxygen.
b. sodium.
c. chlorine.
d. fluorine.

- A 3. One-half the distance between the nuclei of identical atoms that are bonded together is called the
a. atomic radius.
b. atomic diameter.
c. atomic volume.
d. electron cloud.

- C 4. A positive ion is known as a(n)
a. ionic radius.
b. valence electron.
c. cation.
d. anion.

- D 5. A negative ion is known as a(n)
a. ionic radius.
b. valence electron.
c. cation.
d. anion.

- A 6. Across a period in the periodic table, atomic radii
a. gradually decrease.
b. gradually decrease, then sharply increase.
c. gradually increase.
d. gradually increase, then sharply decrease.

- C 7. Which of the following elements has the largest atomic radius?
a. Fluorine
b. Bromine
c. Iodine
d. Chlorine

- A 8. Which of the following elements has the largest atomic radius?
a. Chromium
b. Manganese
c. Zinc
d. Iron

- C 9. Which of the following elements has the largest atomic radius?
a. Calcium
b. Magnesium
c. Barium
d. Strontium

- D 10. Which of the following elements has the largest ionization energy?
a. Hydrogen
b. Helium
c. Lithium
d. Beryllium

- A 11. Which of the following elements has the largest ionization energy?
a. Phosphorous
b. Arsenic
c. Selenium
d. Carbon

- B 12. Which of the following elements has the largest ionization energy?
a. Calcium
b. Magnesium
c. Barium
d. Strontium

Atomic Radius

1. Atoms get **larger** going down a group because adding new shells/energy levels

2. Atoms get **smaller** going across a period because not adding energy levels but $\uparrow Z_{eff}$

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Patrick Vollmar



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