

Write the general stoichiometric equation, and under each term indicate the number of moles of that component:



Initially the total number of moles present in the system is

$$N_0 = N_{A0} + N_{B0} + \cdots + N_{R0} + N_{S0} + \cdots + N_{\text{inert}}$$

but at time t it is

$$N = N_0 + x(r + s + \cdots - a - b - \cdots) = N_0 + x \Delta n \quad (3)$$

where

$$\Delta n = r + s + \cdots - a - b - \cdots$$

Assuming that the ideal gas law holds, we may write for any reactant, say A in the system of volume V

$$C_A = \frac{p_A}{RT} = \frac{N_A}{V} = \frac{N_{A0} - ax}{V} \quad (4)$$

Combining Eqs. 3 and 4 we obtain

$$C_A = \frac{N_{A0}}{V} - \frac{a}{\Delta n} \frac{N - N_0}{V}$$

or

$$p_A = C_A RT = p_{A0} - \frac{a}{\Delta n} (\pi - \pi_0) \quad (5)$$

Equation 5 gives the concentration or partial pressure of reactant A as a function of the total pressure π at time t , initial partial pressure of A , p_{A0} , and initial total pressure of the system, π_0 .

Similarly, for any product R we can find

$$p_R = C_R RT = p_{R0} + \frac{r}{\Delta n} (\pi - \pi_0) \quad (6)$$

Equations 5 and 6 are the desired relationships between total pressure of the system and the partial pressure of reacting materials.

It should be emphasized that if the precise stoichiometry is not known, or if more than one stoichiometric equation is needed to represent the reaction, then the "total pressure" procedure cannot be used.

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Table of Contents Octave Levenspiel Chapter 1solutions

1. Understanding the eBook Octave Levenspiel Chapter 1solutions
 - The Rise of Digital Reading Octave Levenspiel Chapter 1solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Octave Levenspiel Chapter 1solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Octave Levenspiel Chapter 1solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Octave Levenspiel Chapter 1solutions
 - Personalized Recommendations
 - Octave Levenspiel Chapter 1solutions User Reviews and Ratings
 - Octave Levenspiel Chapter 1solutions and Bestseller Lists
5. Accessing Octave Levenspiel Chapter 1solutions Free and Paid eBooks
 - Octave Levenspiel Chapter 1solutions Public Domain eBooks

- Octave Levenspiel Chapter 1solutions eBook Subscription Services
- Octave Levenspiel Chapter 1solutions Budget-Friendly Options
- 6. Navigating Octave Levenspiel Chapter 1solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Octave Levenspiel Chapter 1solutions Compatibility with Devices
 - Octave Levenspiel Chapter 1solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Octave Levenspiel Chapter 1solutions
 - Highlighting and Note-Taking Octave Levenspiel Chapter 1solutions
 - Interactive Elements Octave Levenspiel Chapter 1solutions
- 8. Staying Engaged with Octave Levenspiel Chapter 1solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Octave Levenspiel Chapter 1solutions
- 9. Balancing eBooks and Physical Books Octave Levenspiel Chapter 1solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Octave Levenspiel Chapter 1solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Octave Levenspiel Chapter 1solutions
 - Setting Reading Goals Octave Levenspiel Chapter 1solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Octave Levenspiel Chapter 1solutions
 - Fact-Checking eBook Content of Octave Levenspiel Chapter 1solutions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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