

Question			Answer	M	Guidance
6	a	i	$\lambda$ is the (de Broglie) wavelength <u>associated with a particle</u> $h$ is (a constant known as) the Planck constant $mv$ is the momentum/mass $\times$ speed of <u>the particle</u>	B1 B1 B1	allow object/electron for particle accept velocity for speed accept $m$ and $v$ separately but the particle must appear with one of the quantities to score the mark
		ii	electrons passing through a (thin) sheet of graphite are diffracted producing rings/pattern (on a fluorescent screen)	B1 A1	any suitable situation <b>NOT</b> Au or Al, etc. what is observed + interpretation 1 mark. <b>Do not</b> award the A mark alone unless a plausible situation has been suggested. Young slits type experiments etc score 0/2
		iii 1	$eV = \frac{1}{2} mv^2$ $2eV/m = v^2$ so $v = \sqrt{(2 \times 1.6 \times 10^{-19} \times 5.0 \times 10^5 / 9.1 \times 10^{-31})}$ so $v = 1.3 \times 10^8 \text{ (m s}^{-1}\text{)}$	C1 C1 B1 A0	allow $eV = 8.0 \times 10^{-14} \text{ J}$ for 1 mark allow $\sqrt{(2 \times 8.0 \times 10^{-14} / 9.1 \times 10^{-31})}$ or evidence of correct calculation, e.g. $v = 1.325 \times 10^8$
		iii 2	$\lambda = h/mv = 6.63 \times 10^{-34} / 1.3 \times 10^8 \times 9.1 \times 10^{-31}$ $\lambda = 5.6 \times 10^{-12} \text{ (m)}$	C1 A1	allow $5.5 \times 10^{-12}$ if $1.325 \times 10^8$ is used
		iii 3	$\lambda$ of visible light $5.6 \times 10^{-7} \text{ m}$ so power of 10 = 5 or -5	B1 B1	ect all 2; accept 4 to 7 $\times 10^{-7} \text{ m}$ <b>NOT</b> $10^{-7}$ <b>NOT</b> $10^5$ on the answer line
	b		The photoelectric effect	B1	QWC mark <b>allow</b> exp't description: uv light shone on a zinc plate connected to a gle <b>max 3</b> from 6 marking points
			individual photons are absorbed by individual electrons in the metal surface, i.e. one to one interaction.	B1	
			Only photons with energies above the work function energy will cause photoelectron emission/idea of threshold frequency	B1	
			Hence u-v photons or blue photons will cause photoemission but red photons will not.	B1	
			Number of electrons emitted depends on light intensity	B1	
			A wave model does not explain instantaneous emission of electrons.	B1	
			A wave model does not explain a threshold frequency/wavelength for emission to occur	B1	allow wave model does not explain no emission however bright the light if energy of photon below work function
			<b>Total question 6</b>	<b>16</b>	

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**M Mosston**



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