

Name: _____ Radioactive Decay Practice Date: _____ Period: _____

Background:

- Contains isotopes that decay at **regular** rates in nature
- We use these regular rates to be able to tell how old a rock is by how much is left inside it.
- **Strategy:** Use a chart to determine what percentage is left.
- **Main Ideas:**
 - o Carbon-14 is the only N isotope that is small enough to be used on once living things (plants)- used for fossils less than 70,000 years old
 - o U^{238} and Pb^{210} - used to measure the age of the oldest rocks because their N life is so long
 - o When the N life passes, one element changes to the other. Ex: Carbon changes to Nitrogen.

Radioactive Decay Data

RADIOACTIVE ISOTOPE	DISINTEGRATION	HALF-LIFE (years)	
Carbon-14	$^{14}\text{C} \rightarrow ^{14}\text{N}$	5.7×10^3	
Potassium-40	$^{40}\text{K} \begin{cases} \rightarrow ^{40}\text{Ar} \\ \rightarrow ^{40}\text{Ca} \end{cases}$	1.2×10^9	
Uranium-238	$^{238}\text{U} \rightarrow ^{206}\text{Pb}$	4.5×10^9	
Rubidium-87	$^{87}\text{Rb} \rightarrow ^{87}\text{Sr}$	4.9×10^{10}	

Radioactive Dating Worksheet Answers

**John D. Cutnell, David Young, Kenneth
W. Johnson, Shane Stadler**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the authors' names.

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Table of Contents Radioactive Dating Worksheet Answers

1. Understanding the eBook Radioactive Dating Worksheet Answers
 - The Rise of Digital Reading Radioactive Dating Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Radioactive Dating Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Radioactive Dating Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Radioactive Dating Worksheet Answers
 - Personalized Recommendations
 - Radioactive Dating Worksheet Answers User Reviews and Ratings
 - Radioactive Dating Worksheet Answers and Bestseller Lists
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 - Radioactive Dating Worksheet Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Radioactive Dating Worksheet Answers
 - Highlighting and Note-Taking Radioactive Dating Worksheet Answers
 - Interactive Elements Radioactive Dating Worksheet Answers
 8. Staying Engaged with Radioactive Dating Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Radioactive Dating Worksheet Answers
 9. Balancing eBooks and Physical Books Radioactive Dating Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Radioactive Dating Worksheet Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Radioactive Dating Worksheet Answers
 - Setting Reading Goals Radioactive Dating Worksheet Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Radioactive Dating Worksheet Answers
 - Fact-Checking eBook Content of Radioactive Dating Worksheet Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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