

Physical Science – 2nd Semester – Final Exam Study Guide – (2012-2013)**1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- i. Mass of an object? triple beam balance Basic SI Unit? kg
 - ii. Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - iii. Length of an object? Ruler Basic SI Unit? meter
 - iv. Density of an object? graduated cylinder Basic SI Unit? ml
- b. What is the formula for calculating density? mass/volume
- i. What is the density of a cube that is 15 g and has a volume of 3 cm³? 5g/cm³
 - ii. Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O is less dense than liquid H₂O

2. Chemistry Basics

- a. Element (def.) substance made up of one kind of atom
- b. Atom (def.) basic building block of matter
- c. Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- d. Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atom: Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	<u>1802</u>	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	<u>Billiard Ball Model</u>
J.J. Thomson	<u>1897</u>	Negatively charged electron	<u>Plum Pudding Model</u>
Niels Bohr	<u>1913</u>	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> ; Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another	<u>Bohr Model</u>
Ernest Rutherford	<u>1909</u>	1909 (<u>Based on idea that positive reflects positive</u>) atom mostly empty space; electrons orbit randomly around positively charged nucleus	<u>Planetary Model</u>

4. Periodic Table

- a. The current periodic table is arranged by increasing atomic number.
- b. **Reading the Periodic Table**
- i. An element's properties can be predicted from its location in the periodic table.
 - ii. The atomic number represents the number of protons.
 - iii. The atomic mass represents the number of neutrons and protons.

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Zhe Xu



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Guide to the Evaluation of Educational Experiences in the Armed Services American Council on Education,1978

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Mathematics for Physical Chemistry Robert G. Mortimer,S.M. Blinder,2023-02-20
Mathematics for Physical Chemistry Fifth Edition includes exercises that enable readers to test their understanding and put theory into practice Chapters are constructed around a sequence of mathematical topics progressing gradually into more advanced material before discussing key mathematical skills including the analysis of experimental data and new to this edition complex variables Includes additional new content on Mathematica and its advanced applications Drawing on the experience of its expert authors this book is the ideal supplementary text for practicing chemists and students wanting to sharpen their mathematics skills and understanding of key mathematical concepts for applications across physical chemistry Includes updated coverage of key topics including a review of general algebra and an introduction to group theory Features previews objectives and numerous examples and problems throughout the text to aid learning Provides chemistry specific instruction without the distraction of abstract concepts or theoretical issues in pure mathematics Includes new chapters on complex variables and Mathematica for advanced applications

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