## 2.1.4 & 2.1.5 - Acids and redox

What are standard solutions?	A solution for which its concentration is known accurately.
What 3 things should you ensure when making standard solutions?	Ensure you have clean dry equipment.     Make washings to remove any salt <u>BEFORE MAKING</u> up the solution to the specified amount.     Invert the volumetric flask to ensure everything is dissolved.
How does ammonia react with acid?	<ul> <li>ammonia + acid → ammonium salt</li> <li>NH<sub>3</sub>(aq) + HNO<sub>3</sub> (aq) → NH<sub>4</sub>NO<sub>3</sub>(aq)</li> </ul>
What are strong and weak acids ? Give 3 strong acids and 1 weak acid	<ul> <li>Strong acids dissociate fully to release H' ions in a solution e.g HCl → H' + Cl'</li> <li>Weak acids dissociate partially to release H' ions in a solution – exist in equilibrium</li> <li>Strong acids include: HCl, H<sub>2</sub>SO<sub>4</sub>, and HNO<sub>3</sub>.</li> <li>Weak acids include carboxylic acids such as CH<sub>2</sub>COOH.</li> </ul>
What are strong and weak bases? Give 1 weak base	<ul> <li>Strong bases dissociate fully to release OH ions in a solution</li> <li>Weak bases dissociate partially to release OH ions in a solution – exist in equilibrium e.g. NH<sub>3</sub></li> </ul>
What are diprotic and triprotic acids (with examples) and what does this mean for neutralisation?	<ul> <li>Diprotic acids and triprotic acids donate two and three protons respectively.</li> <li>Examples include: H<sub>2</sub>SO<sub>4</sub> and H<sub>3</sub>PO<sub>4</sub>.</li> <li>They need 2x or 8x the number of moles of base to be neutralized.</li> </ul>

# Ocr Mark Schemes Chemistry May 2013

**L** Cohen

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## Adopting the Melody of Phrase: An Emotional Symphony within Ocr Mark Schemes Chemistry May 2013

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#### **Table of Contents Ocr Mark Schemes Chemistry May 2013**

- 1. Understanding the eBook Ocr Mark Schemes Chemistry May 2013
  - The Rise of Digital Reading Ocr Mark Schemes Chemistry May 2013
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Ocr Mark Schemes Chemistry May 2013
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr Mark Schemes Chemistry May 2013
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Ocr Mark Schemes Chemistry May 2013
  - Personalized Recommendations

- Ocr Mark Schemes Chemistry May 2013 User Reviews and Ratings
- Ocr Mark Schemes Chemistry May 2013 and Bestseller Lists
- 5. Accessing Ocr Mark Schemes Chemistry May 2013 Free and Paid eBooks
  - o Ocr Mark Schemes Chemistry May 2013 Public Domain eBooks
  - o Ocr Mark Schemes Chemistry May 2013 eBook Subscription Services
  - Ocr Mark Schemes Chemistry May 2013 Budget-Friendly Options
- 6. Navigating Ocr Mark Schemes Chemistry May 2013 eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Ocr Mark Schemes Chemistry May 2013 Compatibility with Devices
  - Ocr Mark Schemes Chemistry May 2013 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - $\,\circ\,$  Adjustable Fonts and Text Sizes of Ocr Mark Schemes Chemistry May 2013
  - Highlighting and Note-Taking Ocr Mark Schemes Chemistry May 2013
  - Interactive Elements Ocr Mark Schemes Chemistry May 2013
- 8. Staying Engaged with Ocr Mark Schemes Chemistry May 2013
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Mark Schemes Chemistry May 2013
- 9. Balancing eBooks and Physical Books Ocr Mark Schemes Chemistry May 2013
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Mark Schemes Chemistry May 2013
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Mark Schemes Chemistry May 2013
  - Setting Reading Goals Ocr Mark Schemes Chemistry May 2013
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Mark Schemes Chemistry May 2013
  - Fact-Checking eBook Content of Ocr Mark Schemes Chemistry May 2013

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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