

IONISATION ENERGIES

FIRST IONISATION ENERGY

- Energy required to remove one electron from each atom in 1 mole of gaseous atoms to form 1 mole of gaseous X^+ ions
 - e.g. $Mg(g) \rightarrow Mg^+(g) + e^-$
 - $\Delta H_{1st} = +738 \text{ kJ/mol}$ (endothermic)

SECOND IONISATION ENERGY

- Energy required to remove one electron from each X^+ ion in 1 mole of gaseous ions to form 1 mole of gaseous X^{2+} ions
 - e.g. $Mg^+(g) \rightarrow Mg^{2+}(g) + e^-$
 - $\Delta H_{2nd} = +1461 \text{ kJ/mol}$ (endothermic)
 - Notice X^{2+} is higher than X^+

Ionisation energy depends on:

- The ion
 - More positive charge
 - ☐ more difficult to remove electron
- Distance between nucleus and electron
 - X^{2+} energy is higher than X^+ because the X^{2+} electron is further away from nucleus than second ion
 - ☐ attraction is less and less energy is needed to remove electron
- Shielding effect
 - Shielding electrons are complete shells of electrons between the nucleus and the electron to be removed
 - The inner shells repel the outer electrons, so the more inner shells
 - ☐ stronger shielding and less energy needed to remove electron

ELECTRON ENERGY LEVELS

- Energy levels = electronic shells
- Energy levels given numbers 1 (outermost quantum numbers)
 - e.g. $n=1$ / $n=2$ etc.
- Provides 2 pieces of information
 - Total number of electrons in atom (sum of separate ionisation energies recorded)
 - Number of electron shells present in atom
- Example for Be (Beryllium):

IONISATION ENERGY	kJ/mol	
Be	900	2 electrons away to remove
Be^+	2071	
Be^{2+}	14818	2 electrons hard to remove
Be^{3+}	21006	

- Observations for Be 2 shells
 - Energy level $n=1$ contains low energy electrons (harder to remove)

Ocr June 2014 Chemistry As F321 Paper

Mike Smith



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Table of Contents Ocr June 2014 Chemistry As F321 Paper

1. Understanding the eBook Ocr June 2014 Chemistry As F321 Paper
 - The Rise of Digital Reading Ocr June 2014 Chemistry As F321 Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr June 2014 Chemistry As F321 Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr June 2014 Chemistry As F321 Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr June 2014 Chemistry As F321 Paper
 - Personalized Recommendations
 - Ocr June 2014 Chemistry As F321 Paper User Reviews and Ratings
 - Ocr June 2014 Chemistry As F321 Paper and Bestseller Lists
5. Accessing Ocr June 2014 Chemistry As F321 Paper Free and Paid eBooks
 - Ocr June 2014 Chemistry As F321 Paper Public Domain eBooks

- Ocr June 2014 Chemistry As F321 Paper eBook Subscription Services
- Ocr June 2014 Chemistry As F321 Paper Budget-Friendly Options
- 6. Navigating Ocr June 2014 Chemistry As F321 Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr June 2014 Chemistry As F321 Paper Compatibility with Devices
 - Ocr June 2014 Chemistry As F321 Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr June 2014 Chemistry As F321 Paper
 - Highlighting and Note-Taking Ocr June 2014 Chemistry As F321 Paper
 - Interactive Elements Ocr June 2014 Chemistry As F321 Paper
- 8. Staying Engaged with Ocr June 2014 Chemistry As F321 Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr June 2014 Chemistry As F321 Paper
- 9. Balancing eBooks and Physical Books Ocr June 2014 Chemistry As F321 Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr June 2014 Chemistry As F321 Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr June 2014 Chemistry As F321 Paper
 - Setting Reading Goals Ocr June 2014 Chemistry As F321 Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr June 2014 Chemistry As F321 Paper
 - Fact-Checking eBook Content of Ocr June 2014 Chemistry As F321 Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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