

PERIODIC TABLE TRENDS

Step 1: Complete the squares for each element by adding the atomic number, name, and atomic mass.

Write the atomic number at the top of the square.

Write the atomic name under the symbol.

Write the atomic mass at the bottom of the square.

Step 2: Determine the number of electrons, protons, and neutrons in each statement.

Step 3: Create a Bohr diagram for each element.

Step 4: Draw a Lewis Dot Structure for each element.

Step 5: Show whether the element is a metal, nonmetal, or metalloid by circling the correct response. Also determine if the element is a solid or gas by circling the correct choice.

Step 6: Use the following colors to shade in the square for each element. You should **ONLY** color in the small square in the upper left-hand corner and not the entire card.

Green – Li and Na

Pink – O and S

Blue – Be and Mg

Purple – F and Cl

Orange – B and Al

Red – C and Si

Brown – N and P

Yellow – He, Ne, and Ar

White – H

Step 7: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.

1							2
3	4	5	6	7	8	9	10
11	12	13	14	15	16	17	18

Step 8: Put a title on your table.

Step 9: Make a key for each color with the name of the group for that column.

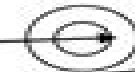
Step 10: Write the column number at the top of each group and the period number on the left side of the table. Be careful, column numbers are not in order from 1-18 and the transition metals, columns 3-12 are not on the table.

Step 11: Draw the red zigzag line that separates the metals from the nonmetals.

Step 12: Answer the questions on the worksheet using the information on your periodic table.

Be

Bohr Diagram



Metal Nonmetal Metalloid

Solid Gas

P = ____

N = ____

E = ____

Be

Section Trends In The Periodic Table Answer Key

Michael Brown



Section Trends In The Periodic Table Answer Key:

Metal Complexes in Aqueous Solutions Arthur E. Martell, Robert D. Hancock, 2013-06-29 Stability constants are fundamental to understanding the behavior of metal ions in aqueous solution. Such understanding is important in a wide variety of areas such as metal ions in biology, biomedical applications, metal ions in the environment, extraction, metallurgy, food chemistry, and metal ions in many industrial processes. In spite of this importance, it appears that many inorganic chemists have lost an appreciation for the importance of stability constants and the thermodynamic aspects of complex formation, with attention focused over the last thirty years on newer areas such as organometallic chemistry. This book is an attempt to show the richness of chemistry that can be revealed by stability constants when measured as part of an overall strategy aimed at understanding the complexing properties of a particular ligand or metal ion. Thus, for example, there are numerous crystal structures of the Li ion with crown ethers. What do these indicate to us about the chemistry of Li with crown ethers? In fact, most of these crystal structures are in a sense misleading in that the Li ion forms no complexes or at best very weak complexes with familiar crown ethers such as 12-crown-4 in any known solvent. Thus, without the stability constants, our understanding of the chemistry of a metal ion with any particular ligand must be regarded as incomplete. In this book, we attempt to show how stability constants can reveal factors in ligand design which could not readily be deduced from any other physical technique.

BIOS Instant Notes in Inorganic Chemistry Tony Cox, 2004-03-01 Instant Notes in Inorganic Chemistry, second edition, has been fully updated and new material added on developments in noble gas chemistry and the synthesis, reactions, and characterization of inorganic compounds. New chapters cover the classification of inorganic reaction types, concentrating on those useful in synthesis; techniques used in characterizing compounds, including elemental analysis, spectroscopic methods, IR, NMR, and structure determination by X-ray crystallography; and the factors involved in choosing appropriate solvents for synthetic reactions. The new edition continues to provide concise coverage of inorganic chemistry at an undergraduate level, offering easy access to all important areas of inorganic chemistry in a format which is ideal for learning and rapid revision.

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Ions in Solution J Burgess,1999-10-01 This outline of the principles and chemical interactions in inorganic solution chemistry delivers a course module in an area of considerable complexity Problems with solutions and tutorial hints to test comprehension have been added as a feature to check readers understanding and assist self study Exercises and projects are also provided to help readers deepen and extend their knowledge and understanding Inorganic solution chemistry is treated thoroughly Emphasis is placed upon NMR UV VIS IR Raman spectroscopy X ray diffraction and such topics as acid base behaviour stability constants and kinetics

Inorganic Chemistry Tina Overton,Jonathan Rourke,Fraser A. Armstrong,2018 Leading the reader from the fundamental principles of inorganic chemistry right through to cutting edge research at the forefront of the subject Inorganic Chemistry Seventh Edition is the ideal course companion for the duration of a student s degree The authors have drawn upon their extensive teaching and research experience to update this text the seventh edition retains the much praised clarity of style and layout from previous editions while offering an enhanced section on expanding our horizons The latest innovative applications of green chemistry have been added to clearly illustrate the real world significance of the subject This edition also sees a greater use of learning features including substantial updates to the problem solving questions additional self tests and walk through explanations which enable students to check their understanding of key concepts and develop problem solving skills Providing comprehensive coverage of inorganic chemistry while placing it in context this text will enable the reader to fully master this important subject Online Resources Inorganic Chemistry Seventh Edition is accompanied by a range of online resources For registered adopters of the text DT Figures marginal structures and tables of data ready to download DT Test bank For students DT Answers to self tests and exercises from the book DT Tables for group theory DT Web links DT Links to interactive structures and other resources on www.chemtube3d.com

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Chemistry in the Community (Enhanced Core Four) American Chemical Society,2006-02-15

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