

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete ionic equation.

Vocabulary

- 1. chemical equation
- 2. balanced equation
- 3. reactant
- 4. coefficient
- 5. product

Part A Completion

Write the coefficient in front of each reactant, product, and catalyst in the chemical equation. Balance the equation. Label the reactants, products, and catalysts.

1. A chemical reaction is usually represented by a chemical _____ 1. _____
 The chemical formula for a chemical compound is _____ 2. _____ The mass
 number for an element is represented by the _____ 3. _____ An element with
 the mass number 40 is _____ 4. _____ A chemical reaction between two or
 more substances is called a _____ 5. _____ reaction. The chemical equation
 represents the change in the number of atoms of each _____ 6. _____ during a reaction. In
 the equation the equation must satisfy the law of conservation of _____ 7. _____
 the chemical equation.

2. The chemical equation for the reaction between hydrogen and oxygen is a
 chemical _____ 8. _____ reaction. The chemical equation is _____ 9. _____
 balanced. The chemical equation is _____ 10. _____ balanced. The chemical equation is
 balanced because the number of atoms of each _____ 11. _____ is the same on both
 sides of the equation. The chemical equation is _____ 12. _____ balanced.

- 1. equation
- 2. formula
- 3. mass number
- 4. element
- 5. chemical
- 6. reaction
- 7. conservation of mass
- 8. chemical
- 9. balanced
- 10. balanced
- 11. element
- 12. balanced

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Léo Pomar

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