ACIDS, BASES AND SALTS

Acids.

- * sowe in darte
- * Changes blue Limus
- * eg: Hydrochloric
 Acid (UC),
 sulphuric Acid
 (H2SO4), Nistric
 Acid (UNO3)

Bases

- · Bitter in taste.
- . shanger red litmus
- ex: sodium hydroxide
 (NaOH), Calcium
 hydroxide (CalOH),),
 Potassium by hydroxide
 (kon)

* some Naturally occurring acids.

- Vinegan
- Онапре
- Lemon
- Tamorind
- · Tomato

- acetic acid.
- litric acid
- Tartavic acid.
- Citric acid.
- Oxalic Doid.

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Didier Musso

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CBSE Class 10 Science Notes , CBSE Class 10 Science Notes Table of Contents Chapter 1 Chemical Reactions and Equations 3 Chapter 2 Acids Bases and Salts 11 Chapter 3 Metals and Non metals 19 Chapter 4 Carbon and its Compounds 28 Chapter 5 Life Processes 38 Chapter 6 Control and Coordination 46 Chapter 7 How do Organisms Reproduce 54 Chapter 8 Heredity 62 Chapter 9 Light Reflection and Refraction 68 Chapter 10 The Human Eye and the Colourful World 76 Chapter 11 Electricity 83 Chapter 12 Magnetic Effects of Electric Current 91 Chapter 13 Our Environment 98 Acid-base Behavior in Aprotic Organic Solvents Marion Maclean Davis,1968 A unified picture of acid base behavior in aprotic organic solvents is presented based on an extensive survey of the literature and experimental results of the author and associates Evidence given to support this picture includes data pertaining to colligative properties of acids bases and salts and also conductance dielectric constants distribution between immiscible solvents and spectral absorption in the infrared visible and ultraviolet The acids upon which attention is centered are proton donor compounds that are measurably ionized in water such as

aliphatic and aromatic carboxylic acids substituted phenols and mineral acids The bases of principal interest are likewise compounds capable of forming ions in water for example aliphatic and aromatic amines and derivatives of guanidine or pyridine The solvents emphasized are hydrocarbons and halohydrocarbons but data for dipolar aprotic solvents for example acetone acetonitrile and nitrobenzene are included Contrasts in acid base behavior and in acidity and basicity scales in aprotic and water like solvents are discussed The role of hydrogen bonding in aprotic solvents is discussed at length Important types of hydrogen bonded structures include chelate rings self associated acids bases and salts hydrogen bonded ion pairs and homo and heteroconjugate cations and anions Examples are given in which hydrogen bonding of these types affects such properties as the absorption spectrum of a salt the catalytic effect of an acid and the accurate location of a titration endpoint Author

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