

**GUIDED PRACTICE PROBLEM 34 (page 307)**

34. Calculate the percent composition of these compounds.

- a. ethane ( $\text{C}_2\text{H}_6$ )
- b. sodium bisulfate ( $\text{NaHSO}_4$ )

a. ethane ( $\text{C}_2\text{H}_6$ )**Analyze****Step 1.** List the knowns and the unknowns.**Knowns**mass of C in one mole ethane =  $2 \times 12.0 \text{ g} = 24.0 \text{ g}$ mass of H in one mole ethane =  $6 \times 1.0 \text{ g} = 6.0 \text{ g}$ molar mass of  $\text{C}_2\text{H}_6 = 24.0 \text{ g} + 6.0 \text{ g} = 30.0 \text{ g}$ **Unknowns**

percent C = ? %

percent H = ? %

Because no masses are given, the percent composition can be determined based on the molar mass of the substance. The percent of an element in a compound is the mass of the element in the compound divided by the mass of the compound. To express the ratio as a percent, the ratio is multiplied by 100%.

**Calculate****Step 2.** Solve for the unknown.

$$\text{percent C} = \frac{24.0 \text{ g C}}{30.0 \text{ g compound}} \times 100\% = 80.0\% \text{ C}$$

$$\text{percent H} = \frac{6.0 \text{ g H}}{30.0 \text{ g compound}} \times 100\% = 20.0\% \text{ H}$$

**Evaluate****Step 3.** Does the result make sense?

The percents of the elements of the compound add up to 100%.

$$80.0\% + 20.0\% = 100\%$$

# Section Chemical Quantities Practice Problem Answers

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