

Name _____ Class _____ Date _____

Assessment

Quiz

Section: Calculating Quantities in Reactions

In the space provided, write the letter of the term or phrase that best answers the question.

- _____ 1. The chemical equation $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$ can be correctly interpreted as
- 1 molecule of $\text{P}_4\text{O}_{10} + 6$ molecules of H_2O form 4 molecules of H_3PO_4 .
 - 1 mol $\text{P}_4\text{O}_{10} + 6$ mol H_2O reacts to form 4 mol H_3PO_4 .
 - 1 g $\text{P}_4\text{O}_{10} + 6$ g H_2O produces 4 g H_3PO_4 .
 - Both (a) and (b)
- _____ 2. Stoichiometry is the branch of chemistry that deals with the _____ between elements and compounds as reactants and products in a chemical reaction.
- bonding
 - energy transfers
 - numerical relationship
 - physical interactions
- _____ 3. How many mole ratios can be correctly obtained from the following chemical equation?
- $$\text{Na}_2\text{S} + \text{Cd}(\text{NO}_3)_2 \rightarrow 2\text{NaNO}_3 + \text{CdS}$$
- 1
 - 4
 - 12
 - 16
- _____ 4. Molar mass can be used to relate
- mass of reactant to moles of reactant.
 - moles of product to mass of product.
 - mass of product to moles of product.
 - All of the above
- _____ 5. In the chemical reaction described by the equation
- $$4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s),$$
- the mole ratio of iron(III) oxide to iron is
- 1:1.
 - 1:2.
 - 2:3.
 - 4:1.

Section Calculating Quantities In Reactions Answer Key

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Holt Chemistry Holt Rinehart & Winston, 2003-01-24 *Ebook: Chemistry: The Molecular Nature of Matter and Change* Silberberg, 2015-01-16 *Ebook Chemistry The Molecular Nature of Matter and Change* Holt Chemistry Ralph Thomas Myers, 2004 *Chemistry3* Andrew Burrows, John Holman, Andrew Parsons, Gwen Pilling, Gareth Price, 2017 *Chemistry3* establishes the fundamental principles of all three strands of chemistry organic inorganic and physical By building on what students have learned at school using carefully worded explanations annotated diagrams and worked examples it presents an approachable introduction to chemistry and its relevance to everyday life **Edexcel A Level Chemistry Student Book 2** Andrew Hunt, Graham Curtis, Graham Hill, 2015-09-25 Exam Board Edexcel Level AS A level Subject Chemistry First Teaching September 2015 First Exam June 2017 Develop and assess your students knowledge and mathematical skills throughout A Level with worked examples practical assessment guidance and differentiated end of topic questions with this Edexcel Year 2 student book Identifies the level of your students understanding with diagnostic questions and a summary of prior knowledge at the start of the Year 1 Student Book Provides support for all 16 required practicals with various activities and questions along with a Practical chapter covering procedural understanding and key ideas related to measurement Mathematical skills are integrated throughout with plenty of worked examples including notes on methods to help explain the strategies for solving each type of problem Offers plenty of practice with Test Yourself Questions to help students assess their understanding and measure progress Encourages further reading and study with short passages of extension material Develops understanding with free online access to Test yourself Answers and an Extended Glossary

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characteristics Chemical sensors are self contained analytical devices that provide real time information on chemical composition A chemical sensor integrates two distinct functions recognition and transduction Such devices are widely used for a variety of applications including clinical analysis environment monitoring and monitoring of industrial processes This text provides an up to date survey of chemical sensor science and technology with a good balance between classical aspects and contemporary trends Topics covered include Structure and properties of recognition materials and reagents including synthetic biological and biomimetic materials microorganisms and whole cells Physicochemical basis of various transduction methods electrical thermal electrochemical optical mechanical and acoustic wave based Auxiliary materials used e g synthetic and natural polymers inorganic materials semiconductors carbon and metallic materials properties and applications of advanced materials particularly nanomaterials in the production of chemical sensors and biosensors Advanced manufacturing methods Sensors obtained by combining particular transduction and recognition methods Mathematical modeling of chemical sensor processes Suitable as a textbook for graduate and final year undergraduate students and also for researchers in chemistry biology physics physiology pharmacology and electronic engineering this book is valuable to anyone interested in the field of chemical sensors and biosensors **Inorganic Chemistry** Mark Weller, Mark T.

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