

Unit 1 Progress Check: MCQ

1. Which of the following numerical expressions gives the number of moles in 5.0 g of CaO?

(A) $5.0 \text{ g} \times 56 \text{ g/mol}$

(B) $\frac{5.0 \text{ g}}{56 \text{ g/mol}}$

(C) $\frac{56 \text{ g/mol}}{5.0 \text{ g}}$

(D) $\frac{1}{5.0 \text{ g}} \times \frac{1}{56 \text{ g/mol}}$



2. In a lab a student is given a 21 g sample of pure Cu metal. Which of the following pieces of information is most useful for determining the number of Cu atoms in the sample? Assume that the pressure and temperature in the lab are 1.0 atm and 25 °C.

(A) The molar mass of Cu

(B) The density of Cu at 25 °C

(C) The volume of the Cu sample

(D) The ratio of the two main isotopes found in pure Cu



3. A 1.0 mol sample of which of the following compounds has the greatest mass?



Novanet Chemistry Answers

Robert Weatherdon



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