

Boyle's Law

- The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
- math
- $P_1 V_1 = P_2 V_2$

Charles' Theory and Boyle's Law

- Expansion of gases is caused by the increase of the gas particles increasing the collisions.
- If the gas is compressible to small the volume is fixed. fixed volume expands the pressure increases with volume.
- Temperature always expands or contracts on the walls of the container.
- Expansion of the gas will increase.

Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. After the volume is reduced, it maintains a volume of 600 ml. above the pressure is 900 mm. What is the volume of the gas at this pressure?

$$P_1 V_1 = P_2 V_2 \quad 800 \text{ ml.} \times 760 \text{ mm.} = 600 \text{ ml.} \times P_2 \quad P_2 = ?$$

$$P_2 = \frac{800 \times 760}{600} = 1013.3 \text{ mm.}$$

Ex. 2. A gas is pressure of 100 mm. and 400 ml. of gas has a volume of 500 ml. If the pressure is increased until the volume is 300 ml. what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad 100 \text{ mm.} \times 400 \text{ ml.} = P_2 \times 300 \text{ ml.} \quad P_2 = ?$$

$$P_2 = 133.3 \text{ mm.}$$

Note Taking Guide Episode 902 Chemistry

Xiang Xie



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